

Thermodynamics Quiz #1 Answer Key

2004

$$(a) \Delta G^\circ = [\Delta_f G^\circ]_{\text{products}} - [\Delta_f G^\circ]_{\text{reactants}} = [-182.47 + 0] - [-162.47 + 0] = -20.0 \text{ kJ/mol}$$

$$K_{\text{eq}} = e^{\frac{\Delta G^\circ}{RT}} = e^{\frac{-20000 \text{ J/mol}}{(8.314 \text{ J/mol}\cdot\text{K})(298 \text{ K})}} = 1.80 \times 10^{-4}$$

$$(b) \Delta H^\circ = [\Delta_f H^\circ]_{\text{products}} - [\Delta_f H^\circ]_{\text{reactants}} = [-202.84 + 0] - [-162.47 + 0] = -37.44 \text{ kJ/mol}$$

$$\Delta S^\circ = \frac{\Delta_f H^\circ_{\text{H}_2\text{O}} - \Delta_f H^\circ_{\text{H}_2\text{O}}}{T} = \frac{22994 \text{ J/mol} - 18884 \text{ J/mol}}{298 \text{ K}} = -13.78 \text{ J/mol}\cdot\text{K}$$

$$\Delta G^\circ = 0 \text{ at eq. } K_{\text{eq}} = \left(\frac{P_{\text{H}_2\text{O}}}{P_{\text{H}_2}} \right)^2 = 1 \text{ at equilibrium}$$

(c)

$$T = \Delta H^\circ / \Delta S^\circ = 22994 \text{ J/mol} / 13.78 \text{ J/mol}\cdot\text{K} = 1669 \text{ K}$$

2003 B

$$(a) \Delta_f H^\circ = \sum \Delta_f H^\circ(\text{products}) - \sum \Delta_f H^\circ(\text{reactants}) = -108.2 - [-137.2 + 2(22)] = -28.8 \text{ kJ/mol}$$

$$(b) K_{\text{eq}} = e^{\frac{\Delta G^\circ}{RT}} = e^{\frac{-28800 \text{ J/mol}}{(8.314 \text{ J/mol}\cdot\text{K})(298 \text{ K})}} = 1.16 \times 10^4$$

$$(c) \Delta_f H^\circ = \frac{\Delta_f H^\circ_{\text{H}_2\text{O}} - \Delta_f H^\circ_{\text{H}_2\text{O}}}{T} = \frac{22994 \text{ J/mol} - 18884 \text{ J/mol}}{298 \text{ K}} = -13.8 \text{ J/mol}\cdot\text{K}$$

(d) Both the standard enthalpy of formation and the standard free energy of formation of acetone = 0.

$$\Delta_f H^\circ_{\text{acetone}} = \frac{\Delta_f H^\circ_{\text{H}_2\text{O}}}{T}$$

$$-313 \text{ J/mol} = 126.8 \text{ J/mol} - 157.8 \text{ J/mol} - 2 \text{ J/mol}$$

$$\Delta_f H^\circ = 131 \text{ J/mol}\cdot\text{K}$$

2004 B

$$(a) \Delta_f H^\circ = \sum \Delta_f H^\circ(\text{products}) - \sum \Delta_f H^\circ(\text{reactants}) = [4(393.5) + 6(215.85) - 2(145.1)] \text{ kJ} = -513.8 \text{ kJ}$$



$$(c) \Delta_f H^\circ_{\text{products}} \text{ and } \Delta_f H^\circ_{\text{reactants}} = [4(57.1) + 4(99.1) + 6(0)] - 2(145.1) = 236.2 - [4(57.1) + 4(99.1)] = -323.8 \text{ kJ/mol}$$

$$(d) \Delta G^\circ = \Delta_f H^\circ - T\Delta_f S^\circ = -513.8 - (298 \text{ K})(-0.523 \text{ kJ/K}) = -347.7 \text{ kJ}$$

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Ensheng Dong

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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