

Cation

Li^+
 Na^+
 K^+

Anion

F^-
 I^-

Compound	Melting Point ($^{\circ}\text{C}$)
LiI	449
KI	686
LiF	845
KF	993



AP Chem 2015

Free response 6

6. A student learns that ionic compounds have significant covalent character when a cation has a polarizing effect on a large anion. As a result, the student hypothesizes that salts composed of small cations and large anions should have relatively low melting points.

- (a) Select two compounds from the table and explain how each supports or refutes the student's hypothesis.
- (b) Identify a compound from the table that can be dissolved in water to produce a basic solution. Write the net ionic equation for the reaction that occurs to cause the solution to be basic.



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