



Name: Chelsey Williams Date: 3/31/22

Student Exploration: Stoichiometry

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, conversion factor, dimensional analysis, formula mass, molar mass, mole, molecular mass, stoichiometry

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? 176 grams
2. It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? 15 sticks of butter

Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The Stoichiometry Gizmo allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:



1. Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

Fe_2O_3 1 CO 3 Fe 2 CO_2 3

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

New equation



New question

How many moles of carbon monoxide (CO) are required to react completely with 1.75 moles of iron (III) oxide (Fe_2O_3)?

Problems solved: 0



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Sunil Tanna

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Table of Contents Answers For Stoichiometry Gizmo Worksheets

1. Understanding the eBook Answers For Stoichiometry Gizmo Worksheets
 - The Rise of Digital Reading Answers For Stoichiometry Gizmo Worksheets
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Stoichiometry Gizmo Worksheets
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Stoichiometry Gizmo Worksheets
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Stoichiometry Gizmo Worksheets
 - Personalized Recommendations
 - Answers For Stoichiometry Gizmo Worksheets User Reviews and Ratings
 - Answers For Stoichiometry Gizmo Worksheets and Bestseller Lists
5. Accessing Answers For Stoichiometry Gizmo Worksheets Free and Paid eBooks
 - Answers For Stoichiometry Gizmo Worksheets Public Domain eBooks
 - Answers For Stoichiometry Gizmo Worksheets eBook Subscription Services
 - Answers For Stoichiometry Gizmo Worksheets Budget-Friendly Options

6. Navigating Answers For Stoichiometry Gizmo Worksheets eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Stoichiometry Gizmo Worksheets Compatibility with Devices
 - Answers For Stoichiometry Gizmo Worksheets Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Stoichiometry Gizmo Worksheets
 - Highlighting and Note-Taking Answers For Stoichiometry Gizmo Worksheets
 - Interactive Elements Answers For Stoichiometry Gizmo Worksheets
8. Staying Engaged with Answers For Stoichiometry Gizmo Worksheets
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Stoichiometry Gizmo Worksheets
9. Balancing eBooks and Physical Books Answers For Stoichiometry Gizmo Worksheets
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Stoichiometry Gizmo Worksheets
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers For Stoichiometry Gizmo Worksheets
 - Setting Reading Goals Answers For Stoichiometry Gizmo Worksheets
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers For Stoichiometry Gizmo Worksheets
 - Fact-Checking eBook Content of Answers For Stoichiometry Gizmo Worksheets
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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