

Unit 1: Atomic Structure and Properties

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 - Group 2: Alkaline Earth Metals
 - Group 3-12: Transition Metals
 - Group 17: Halogens
 - Group 18: Noble Gases
 - First row below table is Lanthanides
 - Second row beneath table are Actinides
- Avogadro's number: 6.022×10^{23} pg 66
- $PV = nRT$ pg. 67
 - N: number of moles
 - P: pressure (atm)
 - V: volume (L)
 - T: temperature (K)
 - R: gas constant, $0.0821 \text{ L atm mol}^{-1}\text{K}^{-1}$
- 1 mol gas: 22.4 L pg. 67
- Energy of electrons is quantized pg. 69
 - Only exist at specific energy levels, each specific intervals apart
- Coulomb's Law: equation for calculating amount of energy electron has a certain distance away from nucleus pg. 70
 - $E = k(+q)(-q)/r^2$
 - E: energy
 - K: Coulomb's constant
 - +q: magnitude of the positive charge (nucleus)
 - -q: magnitude of the negative charge (electron)
 - r: distance between the charges
 - Further away, the greater the energy of electron, with it being 0 at infinity
- Bohr Model: model that predicts electron orbit around fixed radius, but is not accurate Pg. 70
 - Electrons absorb electronic radiation to jump to higher energy levels and release it to fall to lower ones
- Ionization energy: the energy required for the binding energies of an electron to be broken Pg. 71
 - Caused when atom is exposed to more electromagnetic radiation than the binding energies of electrons
 - Energy is spent breaking electron free, and leftover is converted to kinetic energy for ejected electron
 - Incoming Radiation Energy = Binding Energy + Kinetic Energy
 - Requires either visible or ultraviolet range of electromagnetic radiation
- Photoelectron Spectra: amount of ionization energy for all electrons ejected from nucleus Pg. 71

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Michael Brown



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