For questions 1 - 7, choose true or false.

- The rate of a given reaction is constant at any specified temperature. False
- A balanced equation does not indicate the mechanism by which a reaction is proceeding.

  True
- The rate constant is the rate of reaction at constant temperature. False
- In the reaction H<sub>2</sub> + I<sub>2</sub> ↔ 2 HI, only two molecules (H<sub>2</sub> and I<sub>2</sub>) react so this reaction is a bimolecular reaction.
- Relatively few reaction mechanisms have been completely characterized. True (Special ingenuity and techniques are needed to detect fast steps in a reaction mechanism.)
- The rate-determining step of a reaction is the step that proceeds fastest. False
- The half-life of a reaction is the time required for half of the original concentration of the limiting reagent to be consumed in the reaction. True
- Factors that determine reaction rates include all of the following except:
  - a. size of reactant particles
  - b. magnitude of the equilibrium constant (Large values of k do not imply large rates.)
  - c. reaction temperature
  - d. concentration of reactants
- For the reaction 2A + B → C, the reaction rate is:
  - a.  $R = k[A]^2 [B]$
  - b. R = k[A][B]
  - c. R = k[A]
  - d. possibly not any of the above written from the overall reaction.)
- 10. For a reaction of the type A + B → 2C, it is found that doubling the amount of A causes the reaction rate to be four times as great but that doubling the amount of B has no apparent effect on the rate. The rate equation that is most correct is:
  - a. R = k[A]<sup>2</sup> (second order in A and zero order in B)
  - b. R = k[A] [B]
  - c. R k[A]
  - d.  $R = \frac{k[C]^2}{[A][B]}$
- 11. Which of the following is the rate equation for a termolecular elementary reaction?
  - a.  $R = k \{A\}_{A}$
  - b.  $R = k [A]^2$
  - c. R = k [A] [B]
  - d. R = k[A][B][C] (three component collision)

# **Ap Chemistry Chapter 12 Test**

Barron's Educational Series, Neil D. Jespersen, Pamela Kerrigan

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