



AP Chem Investigation 6 – Rate of a Reaction Lab



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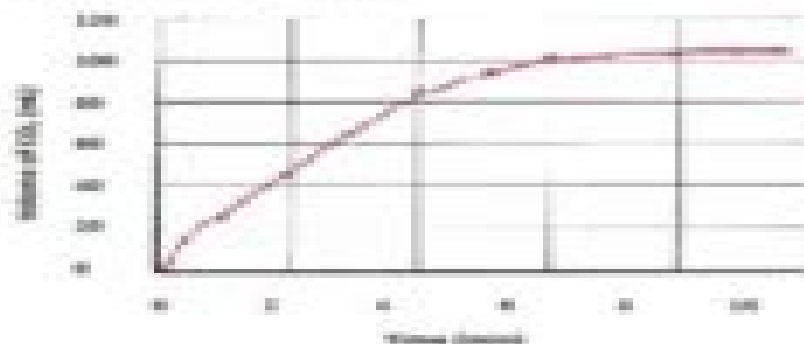
- Statues, monuments, and famous buildings are usually made from limestone or marble, which are minerals that contain large amounts of calcium carbonate, CaCO_3 . Air pollutants have created acid rain that reacts with our statues. In this case, HCl is the rain reacting with CaCO_3 . Write a balanced equation (with predicted products) to describe the reaction between HCl and CaCO_3 . (hint: This is a GULP-powered reaction) $\text{2HCl} + \text{CaCO}_3 \rightarrow \text{CaCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$
- Measure out 20 mL of 6M HCl using a 10 mL graduated cylinder.
- Transfer this acid to a small, clean Erlenmeyer flask.
- Obtain a syringe stopper setup from the front of the room and make sure it fits the flask tightly.
- Position the cutoff valve on the syringe. Turn the valve so that it is parallel with the syringe. This is in the "on" position, so the syringe is open. Give the syringe a twist so be sure it is tightly secured into the stopper system. Push the syringe plunger down all the way to remove as much air as possible.
- Remove the syringe stopper setup and set it aside.
- Measure out as close to 0.10g of CaCO_3 as possible. Record the exact mass that you used in the data table below.
- Add the CaCO_3 to the reaction flask and quickly stopper the flask with the syringe setup. Note the time that you're doing this because you're going to need to take a reading every minute.
- Take a volume measurement on the syringe every minute for the next 10 minutes. To take an accurate reading, you will need to pull up on the syringe plunger and then push it back down. The syringe position should change each time you take a reading. If it doesn't change, ask your teacher for help! You've probably got a leak somewhere!

Mass of CaCO_3 used (grams)	0.10
--------------------------------------	------

Time (minutes)	Volume of CO_2 gas (mL)
0	0.0
1	1.0
2	2.0
3	3.0
4	4.0
5	5.0
6	5.5
7	5.8
8	6.0
9	6.2
10	6.3

Questions:

- Graph the data above in the space below.



- Does the amount of CO_2 produced increase linearly with time, or does it begin to level off as the reaction proceeds?
it levels off
- Explain the shape of the curve based on the rate of the reaction and changes in the concentration of CO_2 over time.
The curve is leveling, so the reaction is slowing its production of CO_2
- As CO_2 is formed, HCl is used up. Explain how the rate of the reaction changes as the HCl concentration changes over time.
As HCl becomes less available, there are fewer collisions between molecules, slowing production to a point it is slower over
- Find the initial rate of formation of CO_2 per minute.
slope of tangent line at (1, 1.0) is 0.1 mL/min
- Find the final rate of formation of CO_2 per minute.
slope of tangent line at (9.5, 6.2) is 0.04 mL/min

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