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AP CHEMISTRY MULTIPLE CHOICE QUESTIONS

Unit 8 – Introduction to Equilibrium



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



Reversed
+ Doubled

(A) 95²

(B) 95

(C) (95)^{1/2}(D) $\frac{1}{95}$ (E) $\frac{1}{(95)^2}$ 

$C_2H_2(g)$ is reduced by $H_2(g)$ in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of $H_2(g)$
- II. Changes in the particle size of the platinum catalyst
- III. Changes in the temperature of the reaction system

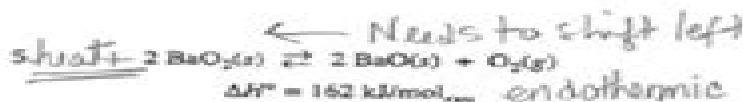
(A) III only

(B) I and II only

(C) I and III only

(D) II and III only

(E) I, II, and III



A sealed rigid vessel contains $BaO_2(s)$ in equilibrium with $BaO(s)$ and $O_2(g)$ as represented by the equation above. Which of the following changes will increase the amount of $BaO_2(s)$ in the vessel?

(A) Removing a small amount of $O_2(g)$ $R \times$ (B) Removing a small amount of $BaO(s)$ X (C) Adding He gas to the vessel X (D) Lowering the temperature L

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products?

increase P°



shift towards fewer
+ make gas



When 4.00 mol each of $X(g)$ and $Y(g)$ are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of $Z(g)$ is produced. What is the value of the equilibrium constant, K_c ?

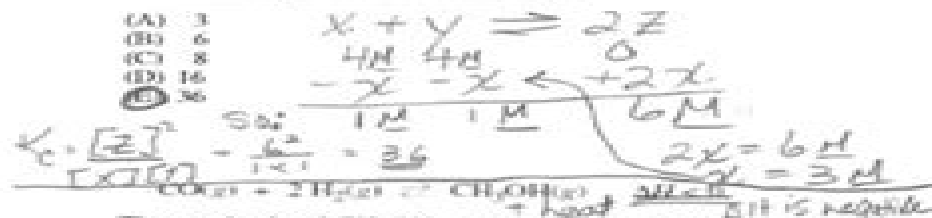
(A) 3

(B) 6

(C) 8

(D) 16

(E) 36



6. The synthesis of $CH_3OH(l)$ from $CO_2(g)$ and $H_2(g)$ is represented by the equation above. The value of K_c for the reaction at 483 K is 14.5.

increase heat shifts left, decrease K_c

Which of the following explains the effect on the equilibrium constant, K_c , when the temperature of the reaction system is increased to 650 K?

(A) K_c will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.(B) K_c will increase because there are more reactant molecules than product molecules.(C) K_c will decrease because the reaction is exothermic.(D) K_c is constant and will not change.

Ap Chemistry Multiple Choice 2008 Answers

**Simone Brogi, Teodorico Castro
Ramalho, José L. Medina-Franco, Kamil
Kuca, Marian Valko**

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AP Chemistry For Dummies Peter J. Mikulecky, Michelle Rose Gilman, Kate Brutlag, 2008-11-13 A practical and hands on guide for learning the practical science of AP chemistry and preparing for the AP chem exam Gearing up for the AP Chemistry exam AP Chemistry For Dummies is packed with all the resources and help you need to do your very best Focused on the chemistry concepts and problems the College Board wants you to know this AP Chemistry study guide gives you winning test taking tips multiple choice strategies and topic guidelines as well as great advice on optimizing your study time and hitting the top of your game on test day This user friendly guide helps you prepare without perspiration by developing a pre test plan organizing your study time and getting the most out of your AP course You ll get help understanding atomic structure and bonding grasping atomic geometry understanding how colliding particles produce states and so much more To provide students with hands on experience AP chemistry courses include extensive labwork as part of the standard curriculum This is why the book dedicates a chapter to providing a brief review of common laboratory equipment and techniques and another to a complete survey of recommended AP chemistry experiments Two full length practice exams help you build your confidence get comfortable with test formats identify your strengths and weaknesses and focus your studies You ll discover how to Create and follow a pretest plan Understand everything you must know about the exam Develop a multiple choice strategy Figure out displacement combustion and acid base reactions Get familiar with stoichiometry Describe patterns and predict properties Get a handle on organic chemistry nomenclature Know your way around laboratory concepts tasks equipment and safety Analyze laboratory data Use practice exams to maximize your score Additionally you ll have a chance to brush up on the math skills that will help you on the exam learn the critical types of chemistry problems and become familiar with the annoying exceptions to chemistry rules Get your own copy of AP Chemistry For Dummies to build your confidence and test taking know how so you can ace that exam

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Subcommittee on Energy and Environment, 2008 **Developing and Validating Test Items** Thomas M. Haladyna, Michael C. Rodriguez, 2013-07-18 Since test items are the building blocks of any test, learning how to develop and validate test items has always been critical to the teaching/learning process. As they grow in importance and use, testing programs increasingly supplement the use of selected response multiple choice items with constructed response formats. This trend is expected to continue. As a result, a new item writing book is needed, one that provides comprehensive coverage of both types of items and of the validity theory underlying them. This book is an outgrowth of the author's previous book *Developing and Validating Multiple Choice Test Items* (3e Haladyna 2004). That book achieved distinction as the leading source of guidance on creating and validating selected response test items. Like its predecessor, the content of this new book is based on both an extensive review of the literature and on its author's long experience in the testing field. It is very timely in this era of burgeoning testing programs, especially when these items are delivered in a computer-based environment. Key features include:
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Strategies for training item writers and developing large numbers of items using algorithms and other item generating methods are also presented.
Based on Theory and Research: A comprehensive review and synthesis of existing research runs throughout the book and complements the expertise of its authors. Library Literature & Information Science, 2008 An index to library and information science literature **Mass Transfer** A. P.

SINHA, PARAMESWAR DE, 2012-05-09 This book introduces the fundamental principles of the mass transfer phenomenon and its diverse applications in process industry. It covers the full spectrum of techniques for chemical separations and extraction. Beginning with molecular diffusion in gases, liquids, and solids within a single phase, the mechanism of interphase mass transfer is explained with the help of several theories. The separation operations are explained comprehensively in two distinct ways: stage-wise contact and continuous differential contact. The primary design requirements of gas-liquid equipment are discussed. The book provides a detailed discussion on all individual gas-liquid, liquid-liquid, liquid-solid, gas and solid-liquid separation processes. The students are also exposed to the underlying principles of the membrane-based separation processes. The book is replete with real applications of separation processes and equipment. Problems are worked out in each

chapter Besides problems with answers short questions multiple choice questions with answers are given at the end of each chapter The text is intended for a course on mass transfer transport and separation processes prescribed for the undergraduate and postgraduate students of chemical engineering [Johnstone Triangle](#) Norman Reid,2021-01-18

Chemistry is often seen as a difficult subject to understand This book focusses on the triangle model that Alex H Johnstone developed in the early 1980s Originally conceived in the context of making chemistry more accessible to a wider range of learners the model has been applied in almost every area of education in chemistry at all stages of learning In looking at why chemistry is difficult there are two central questions Firstly does the problem relate to the nature of chemistry and secondly does it relate to the way humans gain understanding Both were found to be important and the answers to the two question were found to be connected The triangle model arose from sustained research into human learning The central finding from research is the critical role of working memory and the model rationalises so much evidence from chemistry education research as well as the repeated experiences of teachers of chemistry at all levels In order to understand chemistry it is essential to develop sound mental models of molecular reality It generates major implications for the way a chemistry curriculum should be constructed and the processes of teaching and learning in chemistry when the goal is focussed on understanding the key ideas Some of these implications are developed and pointers offered to more successful ways forward The power of the Johnstone Triangle lies in the way it offers clear directions for all involved in chemistry education It is hoped that this book will prove helpful to all involved in sharing the exciting story of the way humans have come to understand the molecular world one of the great examples of great human endeavour *Choice* ,2009 [In Silico Methods for Drug Design and Discovery](#) Simone Brogi,Teodorico Castro Ramalho,José L. Medina-Franco,Kamil Kuca,Marian Valko,2020-10-09 This eBook is a collection of articles from a Frontiers Research Topic Frontiers Research Topics are very popular trademarks of the Frontiers Journals Series they are collections of at least ten articles all centered on a particular subject With their unique mix of varied contributions from Original Research to Review Articles Frontiers Research Topics unify the most influential researchers the latest key findings and historical advances in a hot research area Find out more on how to host your own Frontiers Research Topic or contribute to one as an author by contacting the Frontiers Editorial Office [frontiersin.org](#) about contact

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