

AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

How much sulfur (in grams) is produced from 21.5 grams of O_2 ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$\begin{array}{c} 21.5 \text{ g O}_2 \left| \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \right| \frac{2 \text{ SO}_2}{3 \text{ O}_2} \left| \frac{3 \text{ S}}{\text{SO}_2} \right| \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S} \end{array}$$

B

2. Iron is produced from its ore by the reactions:

How many moles of $\text{O}_2(g)$ are needed to produce 4.6 moles of $\text{Fe}(s)$?A) 2.3 mole O_2 B) 3.5 mole O_2 C) 4.6 mole O_2 D) 6.9 mole O_2 E) 13.8 mole O_2

$$\begin{array}{c} 4.6 \text{ mol Fe}(s) \left| \frac{3 \text{ CO}}{2 \text{ Fe}} \right| \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2 \end{array}$$

B

3. Suppose the reaction
- $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$
- is carried out starting with 153 g of
- $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of H_2SO_4 . How much phosphoric acid will be produced?

A) 76.7 g

B) 51.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$\begin{array}{c} 153 \text{ g Ca}_3(\text{PO}_4)_2 \left| \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{284 \text{ g Ca}_3(\text{PO}_4)_2} \right| \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3(\text{PO}_4)_2} = \\ 76.8 \text{ g H}_2\text{SO}_4 \left| \frac{1 \text{ mol H}_2\text{SO}_4}{98.08 \text{ g H}_2\text{SO}_4} \right| \frac{2 \text{ H}_3\text{PO}_4}{3 \text{ H}_2\text{SO}_4} = \end{array}$$

B

- 4.
- SO_2
- reacts with
- H_2S
- as follows:

When 7.50 g of H_2S reacts with 12.75 g of SO_2 , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.C) 0.0216 moles of H_2S remain.D) 1.13 g of H_2S remain.E) SO_2 is the limiting reagent.

$$\begin{array}{c} 7.50 \text{ g H}_2\text{S} \left| \frac{1 \text{ mol H}_2\text{S}}{34.03 \text{ g H}_2\text{S}} \right| \frac{3 \text{ S}}{2 \text{ H}_2\text{S}} \left| \frac{32 \text{ g S}}{1 \text{ mol S}} \right| = 10.89 \end{array}$$

$$\begin{array}{c} 12.75 \text{ g SO}_2 \left| \frac{1 \text{ mol SO}_2}{64 \text{ g SO}_2} \right| \frac{3 \text{ S}}{1 \text{ SO}_2} \left| \frac{32 \text{ g S}}{1 \text{ mol S}} \right| = 19.13 \end{array}$$

Ap Chemistry Test Answers

Paul Foglino, Princeton Review (Firm)

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2003 Katia Sycara,John Mylopoulos,2003-10-02 These proceedings contain the papers accepted for presentation at the Second International Semantic Web Conference ISWC 2003 held on Sanibel Island Florida U S A October 20 23 2003 Following the success of ISWC 2002 that washeldinSardiniaJune2002 ISWC2003enjoyedagreatlyincreasedinterest in the conference themes The number of submitted papers more than doubled compared with ISWC 2002 to 283 Of those 262 were

submitted to the research track and 21 to the industrial track With rare exceptions each submission was evaluated by three program committee members whose reviews were coordinated by members of the senior program committee This year 49 papers in the research track and 9 papers in the industrial track were accepted The high quality of ISWC 2003 was the result of the joint effort of many people First of all we would like to thank the authors for their high quality submissions and the members of the program committee for their reviewing and review coordination efforts We would like to extend special thanks to Christoph Bussler for chairing the industrial track to Mike Dean for his help with the conference management software the web site and conference publicity and to Massimo Paolucci for helping with the organization of the proceedings and arranging sponsorships

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- Chapter 4: Ap Chemistry Test Answers in Specific Contexts
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