Concepts to know:

Memorize the element symbols (example sodium = Na) States of matter Pure substances (elements and compounds) and mixtures (homogeneous and heterogeneous) Physical and chemical changes 51 units Significant figures Calculations with significant figures Atomic theory of matter History of the atom (cathode ray tube and gold foil experiment) Atomic number and atomic mass Isotopes Average atomic weights Periodic table (metals, nonmetals, metalloids, group names) Formula and molecular weights and molar masses Interconverting masses, moles, and number of particles Percent composition Empirical and molecular formulas Mass spectrometer

Practice problems:

Electrican

1. Give the chemical symbols for the following elements

beard.

- a. Carbon b. cadmium c. silveer. d. sodium e. oxygen f. copper. (C) Call Car Aug. Name the chemical element represented by the following. b. Pb c. NAm. d. Zo er. Feet S. 194
- You are given five pieces of pure aluminum, all different sizes. You measure both the mass and volume of each piece, and plot mass vs. volume on a graph. Assuming your results are accurate, the plot on your graph is

Tribute.

SECURE .

nitrozen

- a. Parabolic b. circular c. a single point d. a straight line e. hyperbolic
- After doing a calculation the display shows the number 3760,289 on the calculator. If the answer is supposed to have three significant figures, the correct way to display the answer would be:
 - a. 3.76 x 10² b 3760. c. 376 d. 376 x 10³ g. 3.76 x 10³

magnesium

- 5. The freezing of water is:
 - a physical change because there is no color change.
 - b. a chemical change because it involves a change in energy
 - a physical change because both ice and liquid water consist solely of H-O
 - d. a chemical change because it does not take place spontaneously
 - e. a chemical change because ice is less dense than liquid water
- 6. A solid sample of a metal is submerged in a large graduated cylinder. The water level in the graduate was 55.25 mL and the final level was 61.00 mL. If the mass of the object was determined to be 23.795 g, what is the density of the metal?
 - a. 4 g/ml. b. 4.14 g/cm² c. 4.13 g/cm² d. 4.1 g/cm³ e. 4.138 g/ml.
- Using the rules for significant figures, calculate the following: (8.790 13)/3.90
 - a. -1.1 b. -1 c. -1.03 d. -1.08 e

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