

OpenQuest 13

Orbital Configuration

Name _____

Date _____

Class _____

Information: Energy of Sublevels

Not all orbitals have the same energy of energy. For example, orbitals in the $3p$ sublevel have more energy than orbitals in the $3s$ sublevel. The following orbitals have sublevels from lowest to highest energy:



Finally, you have information for which orbitals are filled. Here's the orbital configuration for Boron:



Now the 4 orbitals appear in the set of orbitals. These orbitals have the same energy level. For example, the 4 orbitals have higher energy than the 3 orbitals, but the 4 orbitals have a higher energy than the 3 orbitals.

When electrons occupy orbitals, they try to have the lowest possible energy possible. (This is called the **lowest energy principle**.) An electron will occupy a 4 orbital before the 3 orbitals are filled, and a 4 orbital will occupy a 4 orbital before the 3 orbitals are filled. Additionally, the 4 orbitals will be filled. (This is called the **Pauli Exclusion Principle**.) When the electron configuration is filled, they must again be applied to the orbitals—each orbital will have one electron.

Critical Thinking Questions

- a. How many electrons would it take to fill the 4 orbitals? (The 4 orbitals are filled, all the orbitals are filled, and the orbitals are filled.)
b. How many electrons would it take to fill the 4 orbitals? (The 4 orbitals are filled, all the orbitals are filled, and the orbitals are filled.)
- a. How many electrons would it take to fill the 4 orbitals? (The 4 orbitals are filled, all the orbitals are filled, and the orbitals are filled.)
b. How many electrons would it take to fill the 4 orbitals? (The 4 orbitals are filled, all the orbitals are filled, and the orbitals are filled.)

Information: Hund's Rule

Electrons prefer to "split" or "spread". Electrons in the same orbital will have the same energy level. Electrons in the same orbital will have the same energy level. Electrons in the same orbital will have the same energy level.



Electrons in the same orbital will have the same energy level. Electrons in the same orbital will have the same energy level. Electrons in the same orbital will have the same energy level.



Hund's rule is to split the electrons into orbitals. Hund's rule is to split the electrons into orbitals. Hund's rule is to split the electrons into orbitals.

The following are the electron orbital diagrams for the elements, nitrogen and oxygen. Under the diagram's 4 orbitals are filled, and the 4 orbitals are filled.



Critical Thinking Questions

- a. How many "paired" electrons are in a nitrogen atom?
- b. Why does nitrogen's electron configuration is $1s^2, 2s^2, 2p^3$? Why not $1s^2, 2s^2, 2p^4$?
- c. Why does oxygen's electron configuration is $1s^2, 2s^2, 2p^4$? Why not $1s^2, 2s^2, 2p^5$?
- d. Why does nitrogen's electron configuration is $1s^2, 2s^2, 2p^3$?
- e. Why does oxygen's electron configuration is $1s^2, 2s^2, 2p^4$?

Answers For Chemquest

Raymond Chang



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protocolo para la explantaciÓN seguimiento y sustituciÓN de - Jul 14 2023

web la agencia francesa de medicamentos y productos sanitarios a través del sistema de vigilancia de protocolo para la
explantación segu imiento y sustituciÓN de prótesis mamarias p i p aemps 29 de febrero de 2012 3 de 25

cartera comÚn de implantes quirÚrgicos de - Nov 06 2022

web cartera comÚn de implantes quirÚrgicos anexo i orden ssi 1356 2015 de 2 de julio por la que se modifican los anexos ii
iii y vi del real decreto 1030 2006 de 15 de septiembre por el que se establece la cartera de servicios comunes del sistema

nacional de salud y el procedimiento para su