

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of silver nitrate: 5.00 g

Reaction of barium nitrate: 5.00 g

Reaction of silver nitrate solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of silver nitrate solution: 25.0 mL

Procedure:

Reaction and Calculations:

1. Reaction:

$$\text{mole of calcium chloride} = 0.0500 \text{ mol CaCl}_2 \times \frac{1.0}{100.0} = 0.000500 \text{ mol CaCl}_2$$

$$\text{mole of calcium chloride} = 0.0500 \text{ mol CaCl}_2 \times \frac{1.0}{100.0} = 0.000500 \text{ mol CaCl}_2$$

$$\text{mole of calcium chloride produced} = 0.0500 \text{ mol CaCl}_2 \times \frac{1.0}{100.0} = 0.000500 \text{ mol CaCl}_2$$

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2. Reaction equation: $\text{CaCl}_2 + \text{Na}_2\text{CO}_3 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$

3. Reaction equation:

$$\text{mole of CaCl}_2 = 0.0500 \text{ mol CaCl}_2 \times \frac{1.0}{100.0} = 0.000500 \text{ mol CaCl}_2$$

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Answers For Lab 6b Heath Chemistry

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