

Name: Ahmed Faizan Date: 12 May Period:

Specific Heat and Heat Capacity Worksheet

DIRECTIONS: Use $q = (m)(C_p)(\Delta T)$ to solve the following problems. Show all work and units.

Ex: How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$\begin{aligned} q &= ? & \Delta t &= 55 - 22 = 33^\circ\text{C} & q &= m s \Delta t \\ m &= 10\text{g} & s &= 0.9 \text{ J/g}^\circ\text{C} & &= 10 \times 0.9 \times 33 \\ & & & & &= \underline{297 \text{ J}} \end{aligned}$$

1. The temperature of 335 g of water changed from 24.5°C to 26.4°C. How much heat did this sample absorb? C for water = 4.18 J/g°C

$$\begin{aligned} \Delta t &= 26.4 - 24.5 = 2 & q &= m s \Delta t \\ s &= 4.18 \text{ J/g}^\circ\text{C} & &= 335 \times 4.18 \times 2 \\ m &= 335\text{g} & &= \underline{2800.6 \text{ J}} \end{aligned}$$

2. How much heat in kilojoules has to be removed from 225g of water to lower its temperature from 25.0°C to 10.0°C?

$$\begin{aligned} \Delta t &= 25 - 10 = 15 & q &= m s \Delta t \\ m &= 225\text{g} & &= 225 \times 4.18 \times 15 \\ s &= 4.18 \text{ J/g}^\circ\text{C} & &= \underline{14107.5 \text{ J}} \\ 1\text{J} &= 1 \times 10^{-3} \text{ KJ} & \therefore &= \underline{14.1075 \text{ KJ}} \end{aligned}$$

3. To bring 1.0kg of water from 25°C to 99°C takes how much heat input?

$$\begin{aligned} m &= 1\text{kg} = 1000\text{g} & q &= 1000 \times 4.18 \times 74 \\ \Delta t &= 99 - 25 = 74 & q &= \underline{309320 \text{ J}} \\ s &= 4.18 \text{ J/g}^\circ\text{C} & & \end{aligned}$$

An insulated cup contains 75.0g of water at 24.00°C. A 26.00g sample of metal at 82.25°C is added. The final temperature of the water and metal is 28.34°C. What is the specific heat of the metal?

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$$\begin{aligned} (q_{\text{water}} &= -q_{\text{metal}}) & (C = m s) &\rightarrow s = \frac{C}{m} \\ q_{\text{water}} &= 75 \times 4.18 \times (28.34 - 24) & q &= m s \Delta t \\ &= 1252 & 1252 &= 26 \times s \times (82.25 - 28.34) \\ & & & 1252 = 26 \times 53.91 \times s \\ & & & s = \underline{0.893 \text{ J/g}^\circ\text{C}} \end{aligned}$$

5. A calorimeter has a heat capacity of 1265 J/°C. A reaction causes the temperature of the calorimeter to change from 22.34°C to 25.12°C. How many joules of heat were released in this process?

$$\begin{aligned} C &= 1265 \text{ J/}^\circ\text{C} \\ \Delta t &= 25.12 - 22.34 = 2.78 \\ q &= ? \\ q &= C \Delta t \\ &= 1265 \times 2.78 \\ &= \underline{3516.7 \text{ J}} \end{aligned}$$

Answers Of Specific Heat

Mocktime Publication



Answers Of Specific Heat:

Answers to Questions Set at Examinations Great Britain. Education Department. Department of Science and Art, 1889
Answers to questions set at the science and art examinations Science and art department, 1887 Matriculation model answers: heat and light, magnetism and electricity. Being the papers from June 1888 to Jan. 1891, with answers by H.E. Schmitz and R.W. Stewart London univ, exam. papers, 1891 *Physics with Answers* Andrew R. King, Oded Regev, 1997-05-28 This book contains 500 problems covering all of introductory physics along with clear step by step solutions to each problem **Numerical Physics With Chapterwise Question - Answers Class XI by D. C. Upadhyay, Dr. J. P. Goel, Er. Meera Goyal** D. C. Upadhyay, Dr. J. P. Goel, Er. Meera Goyal, 2020-06-27 Unit 1 Physical World and Measurement 1 Systems of Units and Measurement 2 Significant Figures and Error Analysis 3 Dimensional Analysis Unit 2 Kinematics 4 Motion in a Straight Line 5 Vector Analysis 6 Motion in a Plane Unit 3 Laws of Motion 7 Newton's Laws of Motion 8 Friction 9 Uniform Circular Motion Unit 4 Work Energy and Power 10 Work Energy and Power Unit 5 Motion of System of Particles and Rigid Body 11 Centre of Mass 12 Rotational Motion and Moment of Inertia Unit 6 Gravitation 13 Gravitation Unit 7 Properties of Bulk Matter 14 Elasticity 15 Pressure of Fluids 16 Viscosity 17 Surface Tension 18 Temperature and Calorimetry 19 Transfer of Heat Unit 8 Thermodynamics 20 First Law of Thermodynamics 21 Second Law of Thermodynamics Unit 9 Behaviour of Perfect Gases and Kinetic Theory of Gases 22 Behaviour of Perfect Gas and Kinetic Theory of Gases Unit 10 Oscillations and Waves 23 Oscillations 24 Speed of Mechanical Waves and Progressive Waves 25 Superposition of Waves Interference and Beats 26 Reflection of Waves Stationary Waves in Stretched Strings and Organ Pipes 27 Doppler's Effect *Basic Concepts of Chemistry* Leo J. Malone, Theodore O. Dolter, 2011-12-27 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter WileyPLUS sold separately from text *Experimental Physics* William Abbott Stone, 1897 **Review Guide for LPN/LVN Pre-entrance Exam** National League for Nursing, Natl League Nursing, 2008-09-12 Providing an overview of the math science and reading comprehension skills necessary for admission to LPN programs this bestselling guide includes multiple choice questions and practice examinations in each of the three test areas **Physics Question Bank IIT JEE Mains - Chapterwise**, Physics Question Bank IIT JEE Mains Table of Contents Chapter 1 Units and Measurements 3 Chapter 2

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