Chem - Covalent Bonding: Questions & Answers

molecules
Ans - seperate groups of atoms with a formula, bonded together

intermolecular bonds ✓ Ans - bonds between molecules (dispersion, dipole-dipole, hydrogen bonding) determines physical properties

intramolecular bonds
Ans - bonds between atoms in a molecule (covalent, ionic eg.)

covalent bonds
Ans - non metals share electrons to complete the octet rule, an atom forms as many bonds as it needs eg. carbon needs 4 electrons so forms 4 bonds

why do non metals share electrons not transfer?

Ans - non metals have a relatively high electronegativity so competing to attract each others electron, shares instead

electron dot diagrams - Ans - lewis diagram shows valence electrons (bonding pairs) and lone pairs

valence structure
✓ Ans - each pair of bonding electrons is shown as a line, still represents dots as lone pairs. 1 line=2 electrons

bonding pairs and lone pairs
Ans - bonding pairs; shared electrons in bond
lone pairs; non bonding pairs

double covalent bond ✓ Ans - rigid, two pairs of electrons (4 electrons) are shared, represented as a double line in valence structures eg. o2

Answers For Chem Think Covalent Bonding

Leallyn Burr Clapp

Answers For Chem Think Covalent Bonding: Chemistry of the Covalent Bond Leallyn Burr Clapp,1957

Unveiling the Magic of Words: A Review of "Answers For Chem Think Covalent Bonding"

In some sort of defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is truly aweinspiring. Enter the realm of "**Answers For Chem Think Covalent Bonding**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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