

Oxidation-Reduction Worksheet

For each reaction below, identify the atom oxidized, the atom reduced, the oxidizing agent, the reducing agent, the oxidation half reaction, the reduction half reaction, and then balance the equation by the method of oxidation-reduction showing all electrons transfers.

- $$\overset{0}{\text{Mg}} + 2\overset{+1}{\text{H}}\overset{-1}{\text{Cl}} \rightarrow \overset{+2}{\text{Mg}}\overset{-2}{\text{Cl}}_2 + \overset{0}{\text{H}}_2$$

$$\overset{0}{\text{Mg}} \xrightarrow{-2e^-} \overset{+2}{\text{Mg}} \quad 2\overset{+1}{\text{H}} \xrightarrow{+2e^-} \overset{0}{\text{H}}_2$$
- $$\overset{0}{2\text{Fe}} + 3\overset{+3}{\text{V}}_2\overset{-2}{\text{O}}_3 \rightarrow \overset{+3}{\text{Fe}}_2\overset{-2}{\text{O}}_3 + 6\overset{+2}{\text{V}}\overset{-2}{\text{O}}$$

$$2\overset{0}{\text{Fe}} \xrightarrow{-6e^-} \overset{+3}{\text{Fe}}_2 \quad 3\overset{+3}{\text{V}}_2 \xrightarrow{+6e^-} 6\overset{+2}{\text{V}}$$
- $$\overset{+7}{\text{KMnO}}_4 + \overset{+5}{\text{KNO}}_2 + 2\overset{+2}{\text{H}}_2\overset{-2}{\text{SO}}_4 \rightarrow \overset{+2}{\text{Mn}}\overset{-2}{\text{SO}}_4 + 2\overset{+2}{\text{H}}_2\text{O} + \overset{+5}{\text{KNO}}_3 + \overset{-2}{\text{K}}_2\overset{-2}{\text{SO}}_4$$
- $$\text{K}_2\text{Cr}_2\text{O}_7 + \text{SnCl}_2 + \text{HCl} \rightarrow \text{CrCl}_3 + \text{SnCl}_4 + \text{KCl} + \text{H}_2\text{O}$$
- $$\text{KMnO}_4 + \text{NaCl} + \text{H}_2\text{SO}_4 \rightarrow \text{Cl}_2 + \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$$
- $$\text{K}_2\text{Cr}_2\text{O}_7 + \text{H}_2\text{O} + \text{S} \rightarrow \text{SO}_2 + \text{KOH} + \text{Cr}_2\text{O}_3$$
- $$\text{KClO}_3 + \text{C}_{12}\text{H}_{22}\text{O}_{11} \rightarrow \text{KCl} + \text{H}_2\text{O} + \text{CO}_2$$
- $$\text{H}_2\text{C}_2\text{O}_4 + \text{K}_2\text{MnO}_4 \rightarrow \text{CO}_2 + \text{K}_2\text{O} + \text{Mn}_2\text{O}_3 + \text{H}_2\text{O}$$
- $$\text{Mn}(\text{NO}_3)_2 + \text{NaBiO}_3 + \text{HNO}_3 \rightarrow \text{HMnO}_4 + \text{Bi}(\text{NO}_3)_3 + \text{NaNO}_3 + \text{H}_2\text{O}$$
- $$\text{H}_2\text{C}_2\text{O}_4 + \text{KMnO}_4 \rightarrow \text{CO}_2 + \text{K}_2\text{O} + \text{Mn}_2\text{O}_3 + \text{H}_2\text{O}$$

Analyzing Oxidation Reduction Reactions Answer Key

Dr. Alok kumar Dash



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CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

Publications of the National Bureau of Standards ... Catalog United States. National Bureau of Standards, 1978

Handbook of Modern Pharmaceutical Analysis Satinder Ahuja, Stephen Scypinski, 2001 This book describes the role modern pharmaceutical analysis plays in the development of new drugs Detailed information is provided as to how the quality of drug products is assured from the point of discovery until the patient uses the drug Coverage includes state of the art topics such as analytics for combinatorial chemistry and high throughput screening formulation development stability studies international regulatory aspects and documentation and future technologies that are likely to impact the field Emphasis is placed on current easy to follow methods that readers can apply in their laboratories No book has effectively replaced the very popular text Pharmaceutical Analysis that was edited in the 1960s by Tak Higuchi This book will fill that gap with an up to date treatment that is both handy and authoritative

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