Balancing redox equations

Example

Balance the equation showing the oxidation of Fe²⁺ ions to Fe³⁺ ions by dichromate ions (Cr₂O₇²⁻) in an acidic medium

$$Fe^{2+} + Cr_2O_7^{2-} \rightarrow Fe^{3+} + Cr^{3+}$$

Step 1: Identify oxidising and reducing agents and write half reactions

Fe²⁺: +2
$$Cr_2O_7^{2-}$$
: $2Cr + 7(-2) = -2$
Fe³⁺: +3 $2Cr = 12$
 Cr^{3+} : +3 $Cr = +6$

$$Fe^{2+} \rightarrow Fe^{3+}$$

Iron loses one electron, i.e. it is oxidised. It donates one electron to dichromate, i.e. it is the reducing agent.

$$Cr_2^{6+} \rightarrow Cr^{3+}$$

Chromium gains three electrons, i.e. it is reduced. It gains its electrons from iron, i.e. it is the oxidising agent.

Half equation: Reduction reaction: Cr26+ + 3e- → Cr3+

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