

SCH4U Unit 4 - Quiz – Equilibrium

Player: _____

Period: _____

Q1: (10)

Q2: (10)

Q3: (10)

Q4: (10)

Q5: (10)

Write the formulas for:

$$K_{eq} = \frac{[products]}{[reactants]} \quad ; \quad K_{eq} = \frac{P_{products}}{P_{reactants}}$$

$$K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{P_{CH_4}^2 P_{O_2}^2} \quad ; \quad K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{P_{CH_4}^2 P_{O_2}^2}$$

gpt (aq) + h₂o (l) ⇌ gpt(s) + h₂g (g) ; gpt(s) ⇌ h₂g (g)

h₂o (l) ⇌ h₂o (g)

Question 1:

1. Hydrogen sulfide, H₂S, is a slightly poisonous gas. At 100°C, an equilibrium mixture contains 0.020 moles of hydrogen, H₂, 0.010 moles of hydrogen sulfide, and an undetermined amount of sulfur, S₂. The reaction is:



If the volume of the mixture is 1.0 L, what amount of S₂ is present at equilibrium?

a. 0.010

b. 0.020 moles of S₂

c. 0.010 moles of S₂

d. 0.020 moles of S₂

e. Determine the amount of S₂ at equilibrium

f. 0.010

What other species have chemical formula consisting of 2 atoms each?

g. 0.010

	2H ₂ (g)	S ₂ (g)	2H ₂ S(g)
Initial (moles)			
Change (moles)			
Final (moles)	0.020	0.010	0.010

f. 0.010 moles of S₂

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Table of Contents Answers For Sch4u Unit 4

1. Understanding the eBook Answers For Sch4u Unit 4
 - The Rise of Digital Reading Answers For Sch4u Unit 4
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Sch4u Unit 4
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Sch4u Unit 4
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Sch4u Unit 4
 - Personalized Recommendations
 - Answers For Sch4u Unit 4 User Reviews and Ratings
 - Answers For Sch4u Unit 4 and Bestseller Lists
5. Accessing Answers For Sch4u Unit 4 Free and Paid eBooks
 - Answers For Sch4u Unit 4 Public Domain eBooks
 - Answers For Sch4u Unit 4 eBook Subscription Services
 - Answers For Sch4u Unit 4 Budget-Friendly Options

6. Navigating Answers For Sch4u Unit 4 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Sch4u Unit 4 Compatibility with Devices
 - Answers For Sch4u Unit 4 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Sch4u Unit 4
 - Highlighting and Note-Taking Answers For Sch4u Unit 4
 - Interactive Elements Answers For Sch4u Unit 4
8. Staying Engaged with Answers For Sch4u Unit 4
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Sch4u Unit 4
9. Balancing eBooks and Physical Books Answers For Sch4u Unit 4
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Sch4u Unit 4
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers For Sch4u Unit 4
 - Setting Reading Goals Answers For Sch4u Unit 4
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers For Sch4u Unit 4
 - Fact-Checking eBook Content of Answers For Sch4u Unit 4
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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