Answer: A.

Explanation: The value of ΔH is negative. Therefore, the reaction is exothermic or heat-releasing.

Answer: A.

**Explanation:** The entire reaction was multiplied by  $\frac{1}{2}$ . Therefore,  $\triangle H$  must also be halved.

3. Answer: C

Explanation: Option C is incorrect because -566 kJ/mol is the ΔH per mole of O<sub>2</sub> and not for 2 moles. Even though options B and D are expressed in grams, they are still correct because 56 g of CO still corresponds to 2 moles of CO, and 32 g of O<sub>2</sub> still corresponds to 1 mole of O<sub>2</sub>. which is consistent with the expected factor.

4. Answer: A

**Explanation**: 
$$\Delta H = (98 g CO_{(g)})(\frac{1 \mod CO_{(g)}}{28 g CO_{(g)}})(\frac{-566 kf}{2 \mod CO_{(g)}}) = -990.5 kf$$

5. Answer: B

**Explanation**: 
$$\Delta H = (1.5 \, mol \, CO_{2(g)})(\frac{-666kJ}{2 \, mol \, CO_{(hg)}}) = -424.5 \, kJ$$



# **Answer Key Unit 6 Thermochemistry**

J Ma

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