



General Chemistry II Equation Sheet

Think about how to set up the problem first, then apply the needed principles and formulas.

Phase Changes

$$\Delta T_b = iK_b m$$

$$\Delta T_f = iK_f m$$

$$\pi = iMRT$$

$$c = kP$$

$$\frac{P_1}{P_2} = \frac{r_1}{r_2}$$

$$P_{\text{solution}} = \chi_{\text{solvent}} P_{\text{solvent}}^{\circ}$$

$$\Delta P = P_{\text{solvent}}^{\circ} - P_{\text{solution}} = \chi_{\text{solute}} P_{\text{solvent}}^{\circ}$$

Cells

$$E = E^{\circ} - \frac{RT}{nF} \ln(Q)$$

$$E = E^{\circ} - \frac{0.0592}{n} \log(Q) \quad (\text{at } 25^{\circ}\text{C})$$

$$\Delta G = -nFE_{\text{cell}}$$

$$\Delta G^{\circ} = -nFE_{\text{cell}}^{\circ}$$

$$E_{\text{cell}}^{\circ} = \frac{RT}{nF} \ln(K)$$

$$E_{\text{cell}}^{\circ} = \frac{0.0592\text{V}}{n} \log(K) \quad (\text{at } 25^{\circ}\text{C})$$

USE ONE:

$$E_{\text{cell}}^{\circ} = E_{\text{cathode(re d.)}}^{\circ} - E_{\text{anode(re d.)}}^{\circ}$$

$$E_{\text{cell}}^{\circ} = E_{\text{cathode(re d.)}}^{\circ} + E_{\text{anode(ox.)}}^{\circ}$$

Thermodynamics

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta S_{\text{surr}} = \frac{-\Delta H_{\text{sys}}}{T}$$

$$\Delta S_{\text{univ}} = \Delta S_{\text{sys}} + \Delta S_{\text{surr}}$$

$$\Delta S_{\text{ran}}^{\circ} = \Sigma S_{\text{products}}^{\circ} - \Sigma S_{\text{reactants}}^{\circ}$$

$$\Delta G_{\text{ran}}^{\circ} = \Sigma \Delta G_{\text{products}}^{\circ} - \Sigma \Delta G_{\text{reactants}}^{\circ}$$

Concentrations

$$\text{molality} = \frac{\text{moles solute}}{\text{mass solvent (kg)}}$$

$$\% \text{ by mass} = \frac{\text{mass solute}}{\text{mass solute} + \text{mass solvent}} \times 100\%$$

$$\text{parts per million} = \frac{\text{mass solute}}{\text{mass solute} + \text{mass solvent}} \times 10^6$$

$$\text{parts per billion} = \frac{\text{mass solute}}{\text{mass solute} + \text{mass solvent}} \times 10^9$$

Acids and Bases

$$K_w = [H_3O^+][OH^-] = K_a \times K_b$$

$$pOH = -\log[OH^-]$$

$$[OH^-] = 10^{-pOH}$$

$$pH + pOH = pK_a + pK_b = 14.00 \quad (\text{at } 25^{\circ}\text{C})$$

$$pH = pK_a + \log \frac{[A^-]}{[HA]}$$

$$pOH = pK_b + \log \frac{[BH^+]}{[B]}$$

$$\% \text{ ionization} = \frac{[H_3O^+]_{\text{eq}}}{[HA]_{\text{in}}} \times 100\%$$

$$pK_a = -\log(K_a)$$

$$pK_b = -\log(K_b)$$

$$K_a = 10^{-pK_a}$$

$$K_b = 10^{-pK_b}$$

Equilibria

$$Q_c = \frac{[C]^c[D]^d}{[A]^a[B]^b}$$

$$K_c = \frac{[C]^c[D]^d}{[A]^a[B]^b}$$

$$Q_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

$$K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

$$K_p = K_c [R \times T]^{\Delta n}$$

$$\Delta G = \Delta G^{\circ} + RT \ln(Q)$$

$$\Delta G^{\circ} = -RT \ln(K)$$

$$\text{rate} = k[A]^x[B]^y$$

$$k = Ae^{-\frac{E_a}{RT}}$$

$$\ln\left(\frac{k_2}{k_1}\right) = \frac{E_a}{R} \left(\frac{1}{T_1} - \frac{1}{T_2}\right)$$

$$\text{half life} = \frac{\ln(2)}{k} \quad (\text{1st order reactions only})$$

$$[A]_t = -kt + [A]_0$$

$$\ln([A]_t) = -kt + \ln([A]_0)$$

$$\frac{1}{[A]_t} = kt + \frac{1}{[A]_0}$$

Answer Sheet General Chemistry Equations

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Answer Sheet General Chemistry Equations:

General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material Chemical Symbolism and Calculations Stanley Wallace Morse, 1953 **Laboratory Manual**

for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures

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Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed

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Hent, John Andraos, 2019-06-26 *Strategies and Solutions to Advanced Organic Reaction Mechanisms A New Perspective on McKillop's Problems* builds upon Alexander Sandy McKillop's popular text *Solutions to McKillop's Advanced Problems in Organic Reaction Mechanisms* providing a unified methodological approach to dealing with problems of organic reaction mechanism. This unique book outlines the logic, experimental insight and problem solving strategy approaches available when dealing with problems of organic reaction mechanism. These valuable methods emphasize a structured and widely applicable approach relevant for both students and experts in the field. By using the methods described, advanced students and researchers alike will be able to tackle problems in organic reaction mechanism from the simple and straight forward to the advanced.

Development of Learning Strategies Within Chemical Education Dr. Terence McIvor, 2015-02-23 The current status of chemistry practical education and laboratory experience for the first year students is not as encouraging and useful as to produce excellent results. Consequently in this context it needs modification to implement a capstone addressing the category on development and testing of chemistry courses or learning units and demonstrating an approach to implementing a multiweek capstone project as a part of the laboratory curriculum.

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CBSE - ICSE Chemistry Part I Chandan Sengupta, First Publication October 2021. Place of Publication Arabinda Nagar Bankura 722101. This workbook will provide an ample scope in getting exposed to the system of acquiring skills and competence related to the understanding of chemistry. It also exposes the student to the concepts of chemistry for enabling the aspirant in acquisition of skills related to chemistry. Some of the worksheets are prepared along with model answers. Some other worksheets are meant for self assessment and evaluation purposes. It is also observed that some of the topics are specific to the referred curriculum. Some other topics are varyingly incorporated in other streams of study. Culmination of more than two streams will enable the fellow student to cope up with the preparatory works meant for Olympiads and other competitive examinations.

A Laboratory Outline of General Chemistry Alexander Smith, 1899 *ISC Most Likely Question Bank Chemistry Class 12 (2022 Exam) - Categorywise & Chapterwise Topics with Latest Reduced Syllabus, Answering Tips & Mind Maps* Oswal, 2021-06-15 Build your self confidence while preparing from Categorywise Chapterwise Most Likely Question Bank Series for Class 12 ISC.

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