



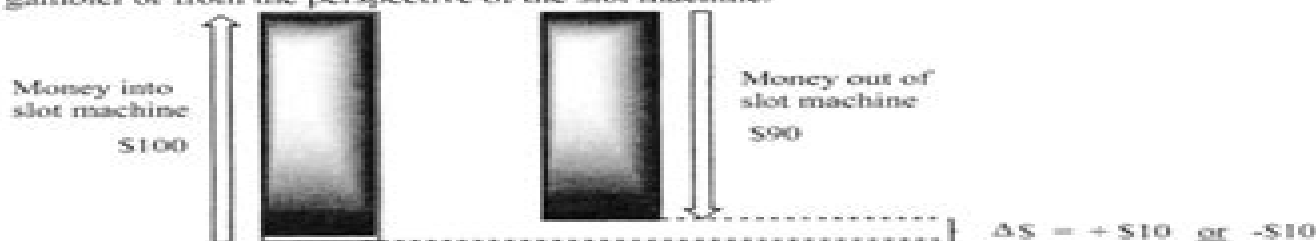
## ChemActivity 6

### Part A: Energy Diagrams

(What is the sign of  $\Delta H_{\text{reaction}}$  for an exothermic reaction?)

#### Model 1A: The Slot Machine

Imagine a gambler sitting at a slot machine. Over the course of an hour the gambler puts \$100 into the machine and the machine pays out \$90. We know a net change ( $\Delta$ ) of ten dollars occurred, but what is the sign of this change in money ( $\Delta S$ ). To determine if  $\Delta S = +10$  or  $-10$  we need to decide if we calculate  $\Delta S$  from the perspective of the gambler or from the perspective of the slot machine.



To translate this analogy into chemistry remember that:

- **Money = energy.**
- **The slot machine = the reaction.**
- **The gambler = the experimenter (you).**
- **All changes ( $\Delta$ ) will be calculated from the perspective of the reaction (a.k.a. the slot machine).**

#### Critical Thinking Questions

1. Using the convention listed above for calculating changes ( $\Delta$ ), what is the sign of  $\Delta S$  in the example in Model 1A?
2. Based on this same convention, what sign must be attached to the \$100 that the gambler put into the slot machine in Model 1A? The sign of the \$90 that comes out is the opposite. **Add the appropriate sign to each value in Model 1A.**
3. On a different day the gambler puts in a total of \$20 and the slot machine pays out a total of \$100. What is the sign and value of  $\Delta S$  for this day?



# Answers Chemactivity 6

**Jiyuan Zhang**



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