

VEGIUM LAB Average Atomic Mass

Objectives

- Determine the weight of each isotopes of the fictitious element "VEGIUM".
- Determine the relative abundances of the isotopes of "VEGIUM".
- Calculate from experimentally data atomic mass of "VEGIUM".

Introduction

Isotopes are atoms of the same atomic number having different masses due to different numbers of neutrons. The atomic mass of an element is a weighted average, taking into account both the mass and relative abundances of each isotope of that element as it occurs in nature. The relative abundances and masses of each atomic particles are measured by the laboratory by the following method as illustrated below. The mass spectrometer separates particles by mass and measures the mass and relative abundance of each type of particle in nature. From these data a weighted average is calculated to determine the atomic mass of the element.

Purpose

In this lab you will carry out experimentally and perform the necessary calculations to determine the mass, mass of isotopes, three isotopes of "VEGIUM" (three isotopes of hydrogen, gases, and black crystal solids). These isotopes are different particles of "VEGIUM" having different masses. Unlike real isotopes, the isotopes will appear particles (gases, gases, and black crystal solids) differ slightly in mass, so you will determine the average mass of each type of isotopic particle. Then you will determine the relative abundance of each isotope, and use the data along with the isotopic masses to calculate the weighted average mass, or atomic mass, of "VEGIUM".

Materials & Equipment

- 3 isotopes of "VEGIUM"
- glass beaker, balance
- weighing boat (that will act as weighing dish)

Experimental Procedure

Carry out the following steps and record your results in your data table. Round all calculations neatly to a maximum of four figures. Report all numbers to the right significant figures.

1. Weigh all three isotopes, three all three gases, three all three black crystal solids. Record your data.
 2. Convert all the gases, three all three gases, three all three black crystal solids. Record your data.
 3. Divide the total mass of each isotope by the number of each isotope to get the average mass of each isotope.
- NOTE: In all cases except the average mass in the table, the numbers in the "total" column are the sum of all numbers, adding a few numbers across each row. It is an exception because it shows each value being summed the fact that there are different

Answers For Vegium Lab

Dennis D. Staley, Edward L. Waterman

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Addison-Wesley Small-scale Chemistry Dennis D. Staley, Edward L. Waterman, 1995

Answers For Vegium Lab Book Review: Unveiling the Power of Words

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