

Experiment 10: Electrochemistry with a Standard Hydrogen Electrode from: *Basic Electrochemistry*, Experiment 10, pp. 45-50

The purpose of this experiment is to determine standard potentials for several other electrodes. In the first part of the experiment, the standard hydrogen electrode (SHE) is used as a reference electrode. The standard potential of the SHE is defined as 0.00 V. The standard potential of the other electrodes is determined by measuring the cell potential of a cell consisting of the SHE and the electrode of interest. The standard potential of the electrode of interest is then calculated from the cell potential. The standard potential of the electrode of interest is then calculated from the cell potential. The standard potential of the electrode of interest is then calculated from the cell potential.

In this experiment, you will determine the standard potentials of several electrodes and calculate the standard potential of the standard hydrogen electrode (SHE) from the standard potential of the other electrodes.

Objectives:

1. To determine the standard potential of the standard hydrogen electrode (SHE) from the standard potential of the other electrodes.
2. To determine the standard potential of the other electrodes from the standard potential of the SHE.
3. To determine the standard potential of the other electrodes from the standard potential of the SHE.
4. To determine the standard potential of the other electrodes from the standard potential of the SHE.

Materials:

- Standard hydrogen electrode (SHE)
- Standard hydrogen electrode (SHE)
- Standard hydrogen electrode (SHE)
- Standard hydrogen electrode (SHE)
- Standard hydrogen electrode (SHE)
- Standard hydrogen electrode (SHE)
- Standard hydrogen electrode (SHE)
- Standard hydrogen electrode (SHE)
- Standard hydrogen electrode (SHE)

Procedure:

1. Prepare the standard hydrogen electrode (SHE) and the other electrodes.
2. Measure the cell potential of the cell consisting of the SHE and the other electrode.
3. Calculate the standard potential of the other electrode from the cell potential.

Results:

1. Calculate the standard potential of the standard hydrogen electrode (SHE) from the standard potential of the other electrodes.
2. Calculate the standard potential of the other electrodes from the standard potential of the SHE.

Discussion:

1. Discuss the results of the experiment and compare them with the standard potentials of the other electrodes.
2. Discuss the results of the experiment and compare them with the standard potentials of the other electrodes.

Answers For Lab 6b Heath Chemistry

Lingjun Ying



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