

SCHAU Unit 4 - Quiz – Equilibrium

Player(s): _____

Period(s): _____

Q1: (10)

Q2: (10)

Q3: (10)

Q4: (10)

Q5: (10)

Relevant Equations:

$$K_c = \frac{[\text{products}]}{[\text{reactants}]} \quad K_p = \frac{(P)^{\text{products}}}{(P)^{\text{reactants}}}$$

$$K_c = \frac{P_{\text{CO}_2}^2 P_{\text{H}_2\text{O}}}{P_{\text{CH}_4}^2 P_{\text{H}_2\text{O}_2}} \quad K_p = \frac{(P_{\text{CO}_2})^2 (P_{\text{H}_2\text{O}})}{(P_{\text{CH}_4})^2 (P_{\text{H}_2\text{O}_2})}$$

gases: $\Delta G_{\text{rxn}} = \Delta G_{\text{f}}(\text{products}) - \Delta G_{\text{f}}(\text{reactants})$ $\Delta G_{\text{rxn}} = -RT \ln K$

$K_p = K_c (RT)^{\Delta n}$

Question 1:

1. Hydrogen sulfide, H_2S , is a slightly poisonous gas. At 1000 K, an equilibrium mixture contains 0.020 moles of hydrogen, H_2 , 0.010 moles of hydrogen sulfide, with an equilibrium pressure of 0.010 atm. The reaction is:



If the volume of an 8.00-L container at 1000 K where a concentration of H_2 is present at equilibrium:

- What is the concentration of H_2 at equilibrium?
- What is the concentration of S_2 at equilibrium?
- What is the concentration of H_2S at equilibrium?
- What is the value of K_c at equilibrium?
- What other species have nonzero chemical formula and are at 1000 K?

	$\text{H}_2(\text{g})$	$\text{S}_2(\text{g})$	$\text{H}_2\text{S}(\text{g})$
Initial (moles)			
Change (moles)			
Final (moles)	0.020	0.010	

What is the value of K_c ?

Answers For Sch4u Unit 4

N Colangelo



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