

MIXED GAS LAWS WORKSHEET - SOLUTIONS

1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ Latm/mol.K})(292 \text{ K})} = 11 \text{ moles of gas}$$

2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of 25°C , what will the pressure of the resulting mixture of gases be? $25^\circ\text{C} = 298 \text{ K}$

$$O_2: P = \frac{nRT}{V} = \frac{(5.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 4.1 \text{ atm}$$

$$N_2: P = \frac{nRT}{V} = \frac{(3.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 2.4 \text{ atm}$$

$$P_{\text{tot}} = P_{O_2} + P_{N_2} = 4.1 \text{ atm} + 2.4 \text{ atm} = 6.5 \text{ atm}$$

Alternately you can substitute total moles of gas (8.0) as the total pressure depends on the total moles of gas – what type of gas is irrelevant

3) A balloon is filled with 35.0 L of helium in the morning when the temperature is 20.0°C . By noon the temperature has risen to 45.0°C . What is the new volume of the balloon?

$$T_1 = 20.0^\circ\text{C} = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^\circ\text{C} = 318 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$$

4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$n = \frac{PV}{RT} = \frac{(190 \text{ atm})(35 \text{ L})}{(0.0821 \text{ Latm/mol.K})(315 \text{ K})} = 260 \text{ moles of gas}$$

5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in $^\circ\text{C}$ of the balloon?

$$T = \frac{PV}{nR} = \frac{(1 \text{ atm})(85 \text{ L})}{(3.5 \text{ mol})(0.0821 \text{ Latm/mol.K})} = 296 \text{ K} = 23^\circ\text{C}$$

6) CaCO_3 decomposes at 1200°C to form CO_2 gas and CaO . If 25 L of CO_2 are collected at 1200°C , what will the volume of this gas be after it cools to 25°C ?

$$T_1 = 1200^\circ\text{C} = 1473 \text{ K}, V_1 = 25 \text{ L}, T_2 = 25^\circ\text{C} = 298 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$

7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at 20.0°C is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is -20.0°C ?

$$P_1 = 1.00 \text{ atm}, V_1 = 4.50 \text{ L}, T_1 = 20.0^\circ\text{C} = 293 \text{ K}, P_2 = 0.600 \text{ atm}, V_2 = ?, T_2 = -20.0^\circ\text{C} = 253 \text{ K}$$

$$V_2 = \frac{P_1 V_1 T_2}{T_1 P_2} = \frac{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}{(293 \text{ K})(0.600 \text{ atm})} = 6.48 \text{ L}$$

Answer Sheet To Mixed Gas Laws

M Mosston

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Answer Sheet To Mixed Gas Laws:

Decoding **Answer Sheet To Mixed Gas Laws**: Revealing the Captivating Potential of Verbal Expression

In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Answer Sheet To Mixed Gas Laws**," a mesmerizing literary creation penned by a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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