Name: __Chelsey Williams_____ Diabe: 3/31/22

Student Exploration: Stoichiometry

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, conversion factor, dimensional analysis, formula mass, molar mass, mole, molecular mass, stoichiometry

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

- A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? 176 grams
- It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? __15 sticks of butter

Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses stoichiometry to determine the amounts of substances involved in chemical reactions. The Stoichiometry Gizmo allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:

 Look at the coefficients (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

New equation

New question

(III) oxide (Fe, Oxid)

Problems solved: 0

Fe₂O₃ (s) + 3CO (p) -- 2Fe (s) + 3CO₂ (p)

How many moles of carbon monoxide (CC) are

required to react completely with 1.75 motes of iron

In a balanced equation, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

Answers For Stoichiometry Gizmo Worksheets

Sunil Tanna

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You Can Do Chemistry Sunil Tanna,2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018

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