

## Practice Test Ch 3

## Stoichiometry

Name: \_\_\_\_\_ Per: \_\_\_\_\_

- Remember, this is practice - Do NOT cheat yourself of finding out what you are capable of doing. Be sure you follow the testing conditions outlined below.
- DO NOT USE A CALCULATOR. AP Chem does not allow the use of a calculator for the MC part of the exam, so it is time to start practicing without one. You may use ONLY a periodic table.
- While you should practice working as fast as possible, it is more important at this point in the course, that you practice without a calculator, even if it slows you down. Look for the "easy math" – common factors and rough estimation – do not do "long division" to try to get exact values. Remember it is a MC test, use the answers
- Mark which questions you would like to "go over" when we get to school in September.

1. Balance the following equation:



The balanced equation shows that 1.00 mole of  $\text{NH}_3$  requires \_\_\_\_\_ mole(s) of  $\text{O}_2$ :

- a. 0.57
  - b. 1.25
  - c. 1.33
  - d. 1.75
  - e. 3.5
2. Write a balanced equation for the combustion of acetaldehyde,  $\text{CH}_3\text{CHO}$ .
- When properly balanced, the equation indicates that \_\_\_\_\_ mole(s) of  $\text{O}_2$  are required for each mole of  $\text{CH}_3\text{CHO}$ .
- a. 1
  - b. 2
  - c. 2.5
  - d. 3
  - e. 5
3. What is the total mass of products formed when 16 grams of  $\text{CH}_4$  is burned with excess oxygen?
- a. 32 g
  - b. 36 g
  - c. 44 g
  - d. 62 g
  - e. 80 g
4. Write a balanced equation for the combustion of propane,  $\text{C}_3\text{H}_8$ .

When balanced, the equation indicates that \_\_\_\_\_ moles of  $\text{O}_2$  are required for each mole of  $\text{C}_3\text{H}_8$ .

- a. 1.5
- b. 3
- c. 3.5
- d. 5
- e. 8

5. Balance the following equation with the SMALLEST WHOLE NUMBER COEFFICIENTS possible. Select the number that is the
- sum
- of the coefficients in the balanced equation:



- a. 3
  - b. 5
  - c. 6
  - d. 7
  - e. 8
6. Calculate the mass of hydrogen formed when 27 g of aluminum reacts with excess hydrochloric acid according to the balanced equation below.
- $$2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$$
- a. 1.5 g
  - b. 2.0 g
  - c. 3.0 g
  - d. 6.0 g
  - e. 12 g
7. How many grams of nitric acid,  $\text{HNO}_3$ , can be prepared from the reaction of 128 g of  $\text{NO}_2$  with 54.0 g  $\text{H}_2\text{O}$  according to the equation below?
- $$3\text{NO}_2 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_3 + \text{NO}$$
- a. 92
  - b. 108
  - c. 126
  - d. 189
  - e. 279
8. Which of the following statements is true?
- I. The molar mass of  $\text{CaCO}_3$  is  $100.1 \text{ g mol}^{-1}$ .
  - II. 50 g of  $\text{CaCO}_3$  contains  $9 \times 10^{23}$  oxygen atoms.
  - III. A 200 g sample of  $\text{CaCO}_3$  contains 2 moles of  $\text{CaCO}_3$ .

- a. I only
- b. II only
- c. III only
- d. I and III only
- e. I, II, and III

# Answer Key To Stoichiometry Test Ch 1

**Peter J. Mikulecky, Michelle Rose  
Gilman, Kate Brutlag**



## **Answer Key To Stoichiometry Test Ch 1:**

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