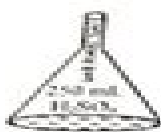


অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পটাসিয়াম কপ্লো, এখানে মোট পটাসিয়াম, আরও কপ্লোয়াম পটাসিয়ামের পটাসিয়াম এবং মনোমের পটাসিয়াম
পটাসিয়াম উত্তর অধ্যায়ভিত্তিক মেট্রিক সিস্টেম। এগুলোর অনুবর্তন করার ক্ষমতা এ অধ্যায়টি থেকে পটাসিয়ামের পটাসিয়াম

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০। ১০০। ১।

- মোলের প্রকরণ কী?
- "কেন্দ্রের পটাসিয়াম লবণের প্রকরণ, 'অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়'— ব্যাখ্যা করো।
- A-পাত্রের বিন্যাসন এলিটের ভর কত নির্ণয় করো।
- প্রকরণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা পরিণত হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL প্রকরণ কোন প্রকরণের 1 mol প্রকরণের প্রকরণ উত্তর প্রকরণের মোলের প্রকরণ করে।

কেন্দ্রের পটাসিয়াম লবণের প্রকরণ অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, অপমাত্রা বৃদ্ধির সাথে সাথে প্রকরণ বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট অপমাত্রার কেন্দ্রের পটাসিয়াম লবণ নিষ্কৃত হয়। তখন প্রকরণ হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কেন্দ্রের পটাসিয়ামে প্রকরণের সময় বিয়োজন প্রক্রিয়াটি অপমাত্রা বৃদ্ধি 32.4°C অপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি অপোৎপাদী হয়। এ কারণে 32.4°C অপমাত্রার উপরে অপমাত্রা বৃদ্ধিজনিত কারণে প্রকরণের হ্রাস ঘটে।

১

অমরা জিনি,
 $w = \text{SMV}$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$

অতএব, প্রকরণ H_2SO_4 এর ভর
 $= 12.25 \text{ g}$

উদীপকের A-পাত্রের অবস্থিতি,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$

প্রকরণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
প্রকরণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$

প্রকরণ H_2SO_4 এর ভর, $w = ?$

১০। পাত্র-A: H_2SO_4 এর মোল সংখ্যা $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

পাত্র-B: Na_2CO_3 এর মোল সংখ্যা $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$

সম্প্রতি বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$

সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 $0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol}$

অবশিষ্ট H_2SO_4 এর পরিমাণ $= (0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

পাত্র-A এর প্রকরণের সাথে পাত্র-B এর প্রকরণের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত প্রকরণের প্রকৃতি হবে অক্সিজেন।

প্রশ্ন ২ এককাল ছাত্র ল্যাবরেটরিতে X-পটাসিয়ামে 10 mg এর প্রকরণ পরিমাণ এবং Y-পটাসিয়ামে 0.5 g পরিমাণ 2% $\frac{w}{v}$ HCl এর প্রকরণ 1000 mL করে।

সি. কে. ১০। ১০০। ১।

- নোড কাকে বলে?
- Cl_2 অপোনার— ব্যাখ্যা করো।
- উদীপকের এলিটটির মোলের ঘনমাত্রা নির্ণয় করো।
- উদীপকের কোন পটাসিয়াম পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

১০। নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেকট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

১০। সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সমান হবে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হবে সমযোজী বন্ধন বিশুদ্ধ ও অপোনার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোনার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অপোনার তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_2 বিশুদ্ধ সমযোজী ও অপোনার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এহুণ ঘটে থাকে।

১০। উদীপকে বর্ণিত এলিটটি হলো 2% $\left(\frac{w}{v}\right)$ HCl।

2% $\left(\frac{w}{v}\right)$ HCl এর অর্থ হলো—

100 mL প্রকরণে HCl এর ভর 2g

∴ 1000 mL " " " " $\left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

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Xiang Xie



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