

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the stoichiometric amount of solid reagents that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reagent 1 (NaCl) Paper: 0.00 g

Reagent 2 (CaCl₂) Paper: 0.00 g

Reaction of NaCl and CaCl₂ solution: 25.0 mL

Reaction of NaCl and CaCl₂ solution: 25.0 mL

Reaction of NaCl and CaCl₂ solution: 25.0 mL

Results:

Observations and Calculations:

1. Calculations:

$$\text{moles of sodium chloride} = 0.00 \text{ mol NaCl} \times \frac{1.0}{100.0} = 0.0000 \text{ mol NaCl}$$

$$\text{moles of calcium chloride} = 0.00 \text{ mol CaCl}_2 \times \frac{1.0}{100.0} = 0.0000 \text{ mol CaCl}_2$$

$$\text{moles of sodium chloride produced} = 0.00 \text{ mol NaCl} \times \frac{1.0}{100.0} = 0.0000 \text{ mol NaCl}$$

$$\text{moles of calcium chloride produced} = 0.00 \text{ mol CaCl}_2 \times \frac{1.0}{100.0} = 0.0000 \text{ mol CaCl}_2$$

2. Chemical Equations: $\text{NaCl} + \text{CaCl}_2 \rightarrow \text{NaCl} + \text{CaCl}_2$

3. Reaction Equation:

$$\text{NaCl} + \text{CaCl}_2 \rightarrow \text{NaCl} + \text{CaCl}_2$$

$$\text{NaCl} + \text{CaCl}_2 \rightarrow \text{NaCl} + \text{CaCl}_2$$

$$\text{NaCl} + \text{CaCl}_2 \rightarrow \text{NaCl} + \text{CaCl}_2$$

Answers For Lab 6b Heath Chemistry

David Baud



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