

Balancing redox equations

Example

Balance the equation showing the oxidation of Fe^{2+} ions to Fe^{3+} ions by dichromate ions ($\text{Cr}_2\text{O}_7^{2-}$) in an acidic medium



Step 1: Identify oxidising and reducing agents and write half reactions



$$2\text{Cr} = 12$$

$$\text{Cr} = +6$$



Iron loses one electron, i.e. it is *oxidised*. It donates one electron to dichromate, i.e. it is the *reducing agent*.

Half equation: Oxidation reaction: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$



Chromium gains three electrons, i.e. it is *reduced*. It gains its electrons from iron, i.e. it is the *oxidising agent*.

Half equation: Reduction reaction: $\text{Cr}_2^{6+} + 3\text{e}^- \rightarrow \text{Cr}^{3+}$

Analyzing Oxidation Reduction Reactions Answers

Yicheng Fang



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Limnological Analysis Robert G. Wetzel, Gene E. Likens, 2013-06-29 Limnological Analyses a classic second thoroughly updated edition consists of a series of carefully designed and tested field and laboratory exercises covering the full scope of limnology It provides the student with a solid foundation in this complex multidisciplinary field of ecology and illustrates modern experimental approaches Among the topics covered by such exercises are major physical components of lakes and streams important mineral nutrients cycling of organic matter benthic fauna primary productivity of phytoplankton quantitative methods in biota analysis diurnal changes experimental manipulation of model ecosystems effects of sewage outfall and other human activities whole ecosystem and community analyses Each exercise is preceded by an introductory section and concludes with questions for the student and a selection of suggested reading Teachers and students of limnology will value Limnological Analyses for its highly structured concise presentation Its research oriented approach encourages active participation

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