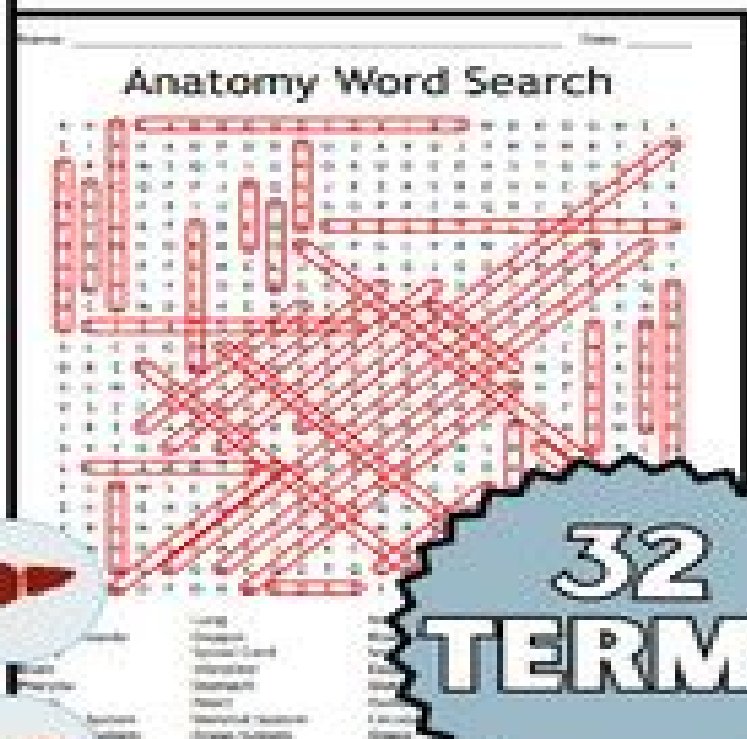
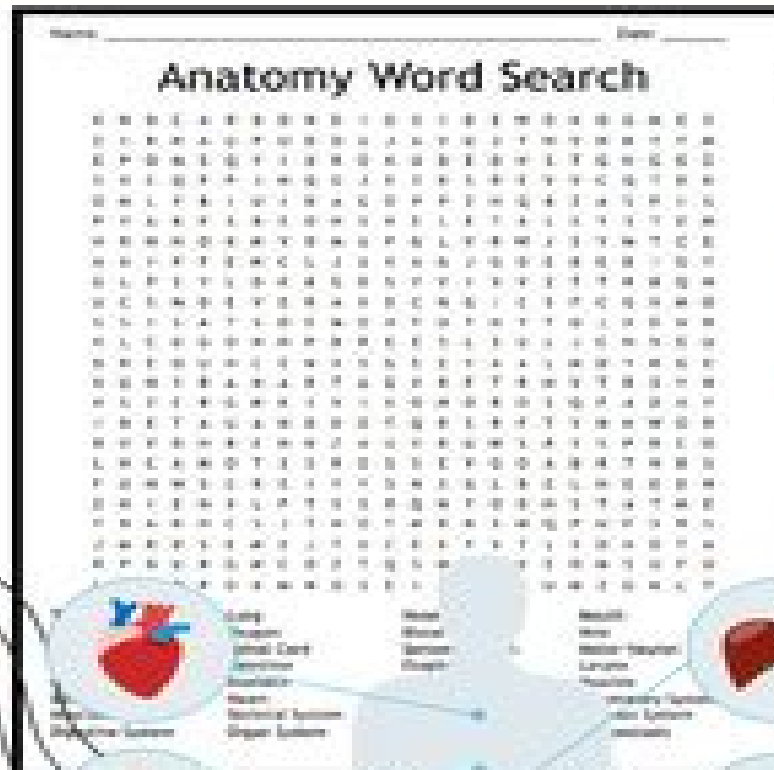


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web limiting and excess reactants 5 13 fill in the table below with the maximum moles of water that can be produced in each container q u indicate which reactant limits the quantity of water produced this is the limiting reactant also show how much of the other reactant the reactant in excess will be left over

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web the limiting reagent concept allows us to calculate amounts of reactants used and products formed in a complete chemical reaction based on the stoichiometric relationships in a balanced chemical equation so far we have considered these relationships on the basis of masses and moles

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web 33mol o₂ 6mol br₂ 11mol o₂ 18mol br₂ identify the limiting reactant s and excess reactant s the limiting reactant is o₂ since it would yield the least amount of product 18 mol br₂ the excess reactant is c₂h₃br₃ since its complete reaction would have yielded up to 54 mol br₂ think about your result

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web step 4 the reactant that produces a smaller amount of product is the limiting reagent mg mg produces less mgo mgo than does o₂ o₂ 3 98 g mgo vs 25 2 g mgo therefore mg is the limiting reagent in this reaction step 5 the reactant that produces a larger amount of product is the excess reagent

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web please save your changes before editing any questions 5 minutes 1 pt viewing model 3 if 4 moles of hydrogen react with 2 moles of oxygen which of these is the excess reactant is leftover at the end and how many moles of it are leftover hydrogen 2 moles leftover hydrogen 1 mole leftover

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web if you start with 14.82 g of Ca(OH)_2 and 16.35 g of H_2SO_4 a determine the limiting reagent b determine the number of moles of H_2O produced c determine the number of grams of CaSO_4 produced d determine the number of grams of excess reagent left 1 make sure the equation is balanced

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