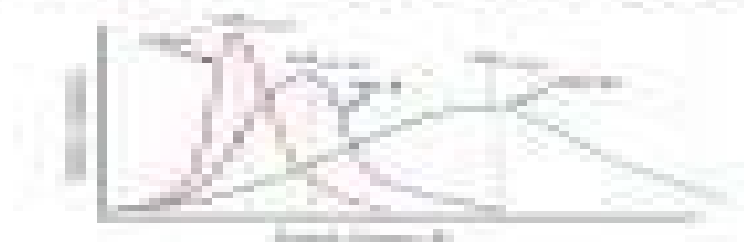


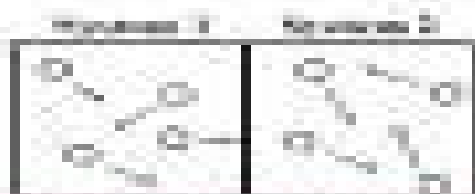
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Free-Energy Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving 100 m/s is 100 J , the first mixing of particles and molecules must not change, instantly increased to 200 J at temperature increase. The Boltzmann distribution curve shows that the average velocity of each of the distribution's molecules must increase. The average temperature increases.

3. Sketch a curve used to represent the relative velocities of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are more spread out than in System 2, so the particles have more kinetic energy. According to the equation $KE = \frac{1}{2}mv^2$, the particles in System 1 have a lower average KE than the particles in System 2, so the particles in System 1 are at a lower temperature than the particles in System 2.
- b. Are the particles in System 1 at a higher or lower temperature than System 2? Justify your answer in terms of molecular collisions.
The particles in System 1 are at a lower temperature than System 2. The particles in System 1 are more spread out than in System 2, so the particles have more kinetic energy. According to the equation $KE = \frac{1}{2}mv^2$, the particles in System 1 have a lower average KE than the particles in System 2, so the particles in System 1 are at a lower temperature than the particles in System 2.

Answer Key Unit 6 Thermochemistry

Lei Shi

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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