Name:	Date:	
	The second second	

Student Exploration: Phase Changes (4b)

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: altitude, boil, boiling point, freeze, freezing point, gas, liquid, melt, melting point, phase, solid

Activity B:	Get the Gizmo ready:	0-000
Temperature and molecular motion	 Click Reset, and select the Micro view. Set Ice volume to 0 cm³. Set Add/remove heat energy to 0 J/s. 	8 6 8 9

Question: Why do phase changes occur?

Compare: Set the Water temperature to 0 °C and click Play. Observe the water molecules. Click Reset, set the Water temperature to 100 °C, and click Play again.

What do you notice?

There is no difference.

- Observe: Click Reset. The mean molecular speed of the water molecules is displayed below the container. Set the Water temperature to 0 °C and Add/remove heat energy to 400 J/s. Click Play.
 - A. How does the mean speed of the water molecules change as they are heated?

The water gets hotter (has more heat energy).

B. Does the mean molecular speed change as much as the temperature as the water heats up? Explain.

Yes, the more the temperature increases the more the molecules of the water speed up (until they hit the boiling point, where they cannot get hotter any longer.)

3. Explain: How is temperature related to the motions of molecules?

The higher the temperature, the faster the molecules move.

4. Observe: Click Reset. Set the Water temperature to 20 °C and the Ice volume to 50 cm³. Set Add/rem ove heat energy to 0 J/s. Click Play. How do the molecules in the liquid interact with the molecules in the solid?

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Ying-Ying Zheng

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