

AP Scoring Guide Chemistry Unit 8 Progress Check MCQ Questions And Answer Keys

Buffer solutions containing Na_2CO_3 and NaHCO_3 range in pH from 10.0 to 11.0. The chemical equation below represents the equilibrium between CO_3^{2-} and H_2O , and the table lists the composition of four different buffer solutions at 25°C .

$$K_b = 2.1$$

Buffer	$[\text{NaHCO}_3]$	$[\text{Na}_2\text{CO}_3]$	pH
1	0.150	0.100	?
2	0.200	0.200	10.32
3	0.100	0.100	10.32
4	0.100	0.200	?

1. Which of the following chemical equilibrium equations best shows what happens in the buffer solutions to minimize the change in pH when a small amount of a strong base is added?

- (A) $\text{H}_3\text{O}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightleftharpoons 2\text{H}_2\text{O}(\text{l})$
- (B) $\text{HCO}_3^-(\text{aq}) + \text{OH}^-(\text{aq}) \rightleftharpoons \text{CO}_3^{2-}(\text{aq}) + \text{H}_2\text{O}(\text{l})$
- (C) $\text{CO}_3^{2-}(\text{aq}) + \text{H}_3\text{O}^+(\text{aq}) \rightleftharpoons \text{HCO}_3^-(\text{aq}) + \text{H}_2\text{O}(\text{l})$
- (D) $\text{CO}_3^{2-}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{HCO}_3^-(\text{aq}) + \text{OH}^-(\text{aq})$

2. Which of the following mathematical expressions can be used to determine the approximate pH of buffer 1?

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John Moore, Richard H. Langley



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