Chemistry

(www.tiwariacademy.com) (Chapter 1)(Solutions) (Intext Questions)

(Class 12)

Question 1.1:

Calculate the mass percentage of benzene (C₆H₆) and carbon tetrachloride (CCl₆) if 22 g of benzene is dissolved in 122 g of carbon tetrachloride.

Answer 1.1:

Mass percentage of $C_6H_6 = \frac{Mass \ of \ C_6H_6}{Total \ mass \ of \ the \ solutions} \times 100 \%$

$$= \frac{Mass\ of\ C_6 H_6}{Mass\ of\ C_6 H_6 + Mass\ of\ CCl_4} \times 100\ \% = \frac{22}{22 + 122} \times 100\ \% = 15.28\ \%$$

Mass percentage of $CCl_4 = \frac{Mass\ of\ CCl_4}{Total\ mass\ of\ the\ solutions} \times 100\ \%$

$$= \frac{Mass\ of\ CCl_4}{Mass\ of\ C_6H_6 + Mass\ of\ CCl_4} \times 100\ \% = \frac{122}{22 + 122} \times 100\ \% = 84.72\ \%$$

Alternatively, Mass percentage of CCI₄ = (100 - 15.28) % = 84.72%

Question 1.2:

Calculate the mole fraction of benzene in solution containing 30% by mass in carbon tetrachloride.

Answer 1.2:

Let the total mass of the solution be 100 g and the mass of benzene be 30 g.

. Mass of carbon tetrachloride =
$$(100 - 30)$$
 g = 70 g

Molar mass of benzene (CsH₆) = $(6 \times 12 + 6 \times 1)$ g mol⁻¹ = 78 g mol⁻¹

. Number of moles of
$$C_6H_6 = \frac{30}{78}$$
 mol = 0.3846 mol

Molar mass of carbon tetrachloride (CCI₄) = $1 \times 12 + 4 \times 355$ = 154 g mol^{-1}

.Number of moles of
$$CCl_4 = \frac{70}{154}$$
 mol = 0.4545 mol

Thus, the mole fraction of CoHo is given as:

Number of moles of C₆H₆ + Number of moles of CCl₄
0.3846

- 0.3846 + 0.4545
- = 0.458

Question 1.3:

Calculate the molarity of each of the following solutions:

Answer 1.3:

Molarity is given by: $Molarity = \frac{Moles \ of \ solution \ in \ litre}{Volume \ of \ solution \ in \ litre}$

- (a) Molar mass of Co (NO₃)₂.6H₂O = $59 + 2(14 + 3 \times 16) + 6 \times 18 = 291 \text{ g mol}^{-1}$
 - .. Moles of Co (NO₃)₂.6H₂O = $\frac{30}{291}$ mol = 0.103 mol

Therefore, molarity =
$$\frac{0.103 \text{ mol}}{4.3 \text{ L}} = 0.023 \text{ M}$$

(b) Number of moles present in 1000 mL of 0.5 M H₂SO₄ = 0.5 mol

. Number of moles present in 30 mL of 0.5 M
$$H_2SO_4 = \frac{0.5\times30}{1000}$$
 mol = 0.015 mol

Therefore, molarity =
$$\frac{0.015 \text{ mol}}{0.5 \text{ L}} = 0.03 \text{ M}$$

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