

1. Vapour pressure of chloroform (CHCl_3) and dichloromethane (CH_2Cl_2) at 25°C are 200 mmHg and 41.5 mmHg respectively. Vapour pressure of the solution obtained by mixing 25.5 g of CHCl_3 and 40 g of CH_2Cl_2 at the same temperature will be :
(Molecular mass of $\text{CHCl}_3 = 119.5 \text{ u}$ and molecular mass of $\text{CH}_2\text{Cl}_2 = 85 \text{ u}$)

- (1) 285.5 mmHg
(2) 173.9 mmHg
(3) 615.0 mmHg
(4) 347.9 mmHg

2. The Gibbs' energy for the decomposition of Al_2O_3 at 500°C is as follows :



$$\Delta_r G^\circ = +960 \text{ kJ mol}^{-1}$$

The potential difference needed for the electrolytic reduction of aluminium oxide (Al_2O_3) at 500°C is at least :

- (1) 5.0 V
(2) 4.5 V
(3) 3.0 V
(4) 2.5 V

3. Four successive members of the first series of the transition metals are listed below. For which one of them the standard potential ($E_{\text{M}^{2+}/\text{M}}^\circ$) value has a positive sign ?

- (1) Fe ($Z = 26$)
(2) Co ($Z = 27$)
(3) Ni ($Z = 28$)
(4) Cu ($Z = 29$)

1. 25°C पर क्लोरोफॉर्म (CHCl_3) तथा डाइक्लोरोमेथेन (CH_2Cl_2) के वाष्प दाब क्रमशः 200 mmHg तथा 41.5 mmHg हैं। 25.5 g CHCl_3 के तथा 40 g CH_2Cl_2 के मिलाने पर प्राप्त विलयन का वाष्प दाब उसी ताप पर होगा,

(CHCl_3 का अणु भार = 119.5 u तथा CH_2Cl_2 का अणु भार = 85 u है)

- (1) 285.5 mmHg
(2) 173.9 mmHg
(3) 615.0 mmHg
(4) 347.9 mmHg

$$\frac{200}{119.5} + \frac{41.5}{85} = \frac{40}{85} + \frac{25.5}{119.5}$$

2. 500°C पर Al_2O_3 के विघटन के लिये गिब्स ऊर्जा निम्नलिखित होती है :



$$\Delta_r G^\circ = +960 \text{ kJ mol}^{-1}$$

500°C पर ऐलुमिनियम ऑक्साइड (Al_2O_3) के विघटन अपचयन के लिये विभव अंतर कम से कम होता है :

- (1) 5.0 V
(2) 4.5 V
(3) 3.0 V
(4) 2.5 V

$$\frac{S}{4} + \frac{1}{2} = \frac{1}{2} \times 1.5$$

3. संक्रमण धातुओं की प्रथम श्रेणी के एक के बाद एक आने वाले चार सदस्य नीचे लिखे जाते हैं। इनमें से किसके लिये स्टैंडर्ड विभव ($E_{\text{M}^{2+}/\text{M}}^\circ$) का मान धनात्मक चिन्ह वाला है ?

- (1) Fe ($Z = 26$)
(2) Co ($Z = 27$)
(3) Ni ($Z = 28$)
(4) Cu ($Z = 29$)

Aipmt 2biology Paper

Michael Seilmaier



Aipmt 2biology Paper:

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