

112. The IUPAC name of



is

- (1) 3, 1-diethyl-4,2,2-dimethylpentane
(2) 4, 4-dimethyl-5, 5-diethylpentane
(3) 5, 5-diethyl-4, 4-dimethylpentane
(4) 3-ethyl-4, 4-dimethylheptane

Ans.
Sol.



The correct answer is 3-ethyl-4, 4-dimethylheptane.

Hence, (4) is correct.

113. Which of the following species exhibits the diamagnetic behaviour?

- (1) O_2^+
(2) O_2
(3) O_2^-
(4) NO

Ans.
Sol.

The correct option is O_2^+

This species has 18 π , which are filled in such a way that all molecular orbitals are fully filled, so diamagnetic.



Hence, (1) is correct.

114. The stability of divalides of Sn, Ge, Si and Pb increases steadily in the sequence

- (1) $\text{GeX}_2 < \text{SnX}_2 < \text{SnX}_2 < \text{PbX}_2$
(2) $\text{SnX}_2 < \text{GeX}_2 < \text{PbX}_2 < \text{SnX}_2$
(3) $\text{SnX}_2 < \text{GeX}_2 < \text{SnX}_2 < \text{PbX}_2$
(4) $\text{PbX}_2 < \text{SnX}_2 < \text{GeX}_2 < \text{SnX}_2$

Ans.
Sol.

Due to inert pair effect, the stability of +2 oxidation state increases as we move down this group.
 $\therefore \text{SnX}_2 < \text{GeX}_2 < \text{SnX}_2 < \text{PbX}_2$

Hence, (3) is correct.

115. Identify the incorrect statement among the following

- (1) Ozone reacts with SO_2 to give SO_3
(2) Silicon reacts with $\text{NaOH}_{\text{conc}}$ in the presence of air to give Na_2SiO_3 and H_2O
(3) Cl_2 reacts with excess of NH_3 to give N_2 and HCl
(4) Br_2 reacts with hot and strong NaOH solution to give NaBr , NaBrO_3 and H_2O

Ans.
Sol.

(4) Br_2 reacts with hot and strong NaOH to give NaBr , NaBrO_2 and H_2O .

Hence, (4) is incorrect statement.

116. The charge/radius ratio of a cation determines its polarizing power. Which one of the following sequences represents the increasing order of the polarizing power of the cationic species, K^+ , Ca^{2+} , Mg^{2+} , Be^{2+} ?

- (1) $\text{Mg}^{2+}, \text{Be}^{2+}, \text{K}^+, \text{Ca}^{2+}$
(2) $\text{Be}^{2+}, \text{K}^+, \text{Ca}^{2+}, \text{Mg}^{2+}$
(3) $\text{K}^+, \text{Ca}^{2+}, \text{Mg}^{2+}, \text{Be}^{2+}$
(4) $\text{Ca}^{2+}, \text{Mg}^{2+}, \text{Be}^{2+}, \text{K}^+$

Ans.
Sol.

(3) Higher the charge/radius ratio, more is the polarizing power.



Hence, (3) is correct.

117. The density (in g mL^{-1}) of a 3.65 M sulphuric acid solution that is 20% H_2SO_4 (Molar mass = 98 g mol^{-1}) by mass will be

- (1) 1.64
(2) 1.88
(3) 1.22
(4) 1.45

Ans.
Sol.

Let the density of solution be d

Molarity of solution given = 3.6

i.e. 1 litre of solution contains 3.6 moles of H_2SO_4

or 1 litre of solution contains 3.6×98 gms of H_2SO_4

Since, the solution is 20% by mass,

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1. Understanding the eBook Aieee Exam Answer For G Code Question
 - The Rise of Digital Reading Aieee Exam Answer For G Code Question
 - Advantages of eBooks Over Traditional Books
2. Identifying Aieee Exam Answer For G Code Question
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Aieee Exam Answer For G Code Question
 - User-Friendly Interface
4. Exploring eBook Recommendations from Aieee Exam Answer For G Code Question
 - Personalized Recommendations
 - Aieee Exam Answer For G Code Question User Reviews and Ratings

- Aieee Exam Answer For G Code Question and Bestseller Lists
- 5. Accessing Aieee Exam Answer For G Code Question Free and Paid eBooks
 - Aieee Exam Answer For G Code Question Public Domain eBooks
 - Aieee Exam Answer For G Code Question eBook Subscription Services
 - Aieee Exam Answer For G Code Question Budget-Friendly Options
- 6. Navigating Aieee Exam Answer For G Code Question eBook Formats
 - ePub, PDF, MOBI, and More
 - Aieee Exam Answer For G Code Question Compatibility with Devices
 - Aieee Exam Answer For G Code Question Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Aieee Exam Answer For G Code Question
 - Highlighting and Note-Taking Aieee Exam Answer For G Code Question
 - Interactive Elements Aieee Exam Answer For G Code Question
- 8. Staying Engaged with Aieee Exam Answer For G Code Question
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Aieee Exam Answer For G Code Question
- 9. Balancing eBooks and Physical Books Aieee Exam Answer For G Code Question
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Aieee Exam Answer For G Code Question
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Aieee Exam Answer For G Code Question
 - Setting Reading Goals Aieee Exam Answer For G Code Question
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Aieee Exam Answer For G Code Question
 - Fact-Checking eBook Content of Aieee Exam Answer For G Code Question
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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