

Prelab

Enzymatic Decomposition and Analysis of Hydrogen Peroxide

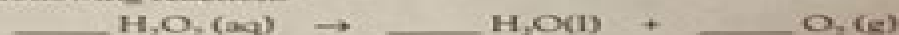
Name _____

Date _____

Section _____

1. What gas is being collected in today's experiment?

2. Balance the following reaction:



3. In today's experiment you determine that you collected 1.33×10^{-3} mols of oxygen. Calculate (show your setup for the conversion!) how many mols of H_2O_2 were in your hydrogen peroxide solution. Note: You will need a conversion factor from the equation you balanced in question 2.

_____ mols of H_2O_2

4. Based on your answer above, how many grams of H_2O_2 were in the hydrogen peroxide solution? (H_2O_2 has a molecular weight of 34.0 g/mol)

_____ grams of H_2O_2

5. Assume the grams of H_2O_2 calculated in question 4 were found in a 6.00 gram sample of a hydrogen peroxide solution. Calculate the mass percent H_2O_2 in the solution.

_____ % H_2O_2

Analysis Of Hydrogen Peroxide Prelab Answers

**Walter G. Cox, NAVAL SURFACE
WEAPONS CENTER SILVER SPRING MD.**

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Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 *Laboratory Manual for Principles of General Chemistry* 11th Edition covers two semesters of a general chemistry laboratory program. The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure, selecting the proper apparatus, employing the proper techniques, systematically analyzing and interpreting the data, and minimizing inherent variables. As a result of good data, a scientific and analytical conclusion is made, which may or may not be right but is certainly consistent with the data. Experiments write textbooks; textbooks don't write experiments. A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered. Further experimentation provides additional cause-effect observations leading to an even better understanding of the experiment. The 11th edition's experiments are informative and challenging while offering a solid foundation for technique, safety, and experimental procedure. The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment. The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit. An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment. Continued emphasis on handling data is supported by the Data Analysis section.

Instructors Manual to Lab Manual Ralph Petrucci, William Harwood, Geoffrey Herring, 2001. *Laboratory Experiments for General Chemistry* Harold R. Hunt, Toby F. Block, 1994. Methods in Biotechnology Seung-Beom Hong, M. Bazlur Rashid, Lory Z. Santiago-Vázquez, 2016-08-01. As rapid advances in biotechnology occur, there is a need for a pedagogical tool to aid current students and laboratory professionals in biotechnological methods. *Methods in Biotechnology* is an invaluable resource for those students and professionals. *Methods in Biotechnology* engages the reader by implementing an active learning approach, providing advanced study questions as well as pre and post lab questions for each lab protocol. These self-directed study sections encourage the reader to not just perform experiments but to engage with the material on a higher level, utilizing critical thinking and troubleshooting skills. This text is broken into three sections based on level: *Methods in Biotechnology*, *Advanced Methods in Biotechnology I*, and *Advanced Methods in Biotechnology II*. Each section contains 14-22 lab exercises with instructor notes in appendices as well as an answer guide as a part of the book companion site. This text will be an excellent resource for both students and laboratory professionals in the biotechnology field.

Experimental Organic Chemistry John C. Gilbert, Stephen F. Martin, 2006. This proven and well-tested laboratory manual for organic chemistry students contains procedures for both miniscale (also known as small scale) and microscale users. This lab manual gives students all the necessary background to enter the laboratory with the knowledge to perform the experiments with confidence. For the microscale labs, experiments were chosen to provide tangible quantities of material which can then be analyzed. Chapters 1-2 introduce students to the

equipment record keeping and safety of the laboratory Chapters 3 6 and 8 are designed to introduce students to laboratory techniques needed to perform all experiments In Chapters 7 and 9 through 20 students are required to use the techniques to synthesize compounds and analyze their properties In Chapter 21 students are introduced to multi step syntheses of organic compounds a practice well known in chemical industry In Chapter 23 students are asked to solve structures of unknown compounds The new chapter 24 introduces a meaningful experiment into the textbook that reflects the increasing emphasis on bioorganic chemistry in the sophomore level organic lecture course This experiment not only gives students the opportunity to accomplish a mechanistically interesting and synthetically important coupling of two amino acids to produce a dipeptide but also provides valuable experience regarding the role of protecting groups in effecting synthetic transformations with multiple functionalized molecules Science Insights ,1999 **Biology** ,1998 *Addison-Wesley Science Insights* ,1996 **Introductory Chemistry** Steven S. Zumdahl,1996 Microscale and Miniscale Organic Chemistry Laboratory Experiments Allen M. Schoffstall,Barbara A. Gaddis,Melvin L. Druelinger,2000 This work offers a comprehensive introductory treatment of the organic laboratory techniques for handling glassware and equipment safety in the laboratory micro and mini scale experimental procedures theory of reactions and techniques applications and spectroscopy **The Software Encyclopedia** ,1985 **Heath Chemistry** ,1993 *The Analysis of Hydrogen Peroxide Solutions* BECCO Chemical Division. Food Machinery and Chemical Corporation,Buffalo Electro-Chemical Company,1954

Analysis of a Hydrogen Peroxide Solution George Deckey,1987-01-01 **The Spectrophotometric Determination of Hydrogen Peroxide** Robert W. Bailey,1956 **A Volumetric Method for Titrimetric Analysis of Hydrogen Peroxide** Walter G. Cox,NAVAL SURFACE WEAPONS CENTER SILVER SPRING MD.,1985 This report documents an oxidation reduction scheme for the volumetric analysis of hydrogen peroxide in the approximate range of 30 percent but it is not limited to this range The purpose of the proposed method is to provide a titrimetric noninstrumental method for the analysis of hydrogen peroxide Keywords Hydrogen Peroxide Potassium Dichromate Ferrous Ammonium Sulfate Hexahydrate Titrimetric Analysis Quantitative Analysis Volumetric Analysis With proper care application of the method described in this technical report provides excellent precision **Hydrogen Peroxide and Cell Signaling, Part B** ,2013-06-29 This new volume of Methods in Enzymology continues the legacy of this premier serial with quality chapters authored by leaders in the field This is the second of three volumes on hydrogen peroxide and cell signaling and includes chapters on such topics as the cellular steady state of H₂O₂ evaluating peroxiredoxin sensitivity towards inactivation by peroxide substrates and peroxiredoxins as preferential targets in H₂O₂ induced signaling Continues the legacy of this premier serial with quality chapters authored by leaders in the field Covers hydrogen peroxide and cell signaling Contains chapters on such topics as the cellular steady state of H₂O₂ evaluating peroxiredoxin sensitivity towards inactivation by peroxide substrates and peroxiredoxins as preferential targets in H₂O₂ induced signaling **The Analysis of Aqueous Solutions of Hydrogen**

Peroxide Charles E. Huckaba,1947 **Determination of Hydrogen Peroxide in Reactor Moderator Solutions by Flow Injection Analysis** ,1991 A flow injection analysis FIA method for the determination of hydrogen peroxide in reactor moderator water was developed and installed at the Savannah River Site SRS Water Quality Laboratory The technique has an analytical range of 0.10 to 2.50 ppm ug/mL with a sampling rate of 40 samples per hour The calibration curve is linear with a correlation coefficient of 0.999 and the precision is excellent with relative standard deviations at the 0.50% level for both 0.10 and 2.50 ppm standards When the automated FIA procedure is compared to the manual method it demonstrates a twenty minute reduction in analysis time per sample and the total liquid waste generated per sample analyzed is reduced by roughly 95 mL 6 refs 10 figs 6 tabs **Theoretical Analysis of the Decomposition of High Strength Hydrogen Peroxide Solutions** Herbert S. Isbin,1947

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