

Analytical Chemistry MCQs: TYBSC SEM III

8. The solubility product K_{sp} of sparingly soluble Ag_2CrO_4 is 4×10^{-12} . The molar solubility of the salt is
(a) 1.0×10^{-4} moles/lit
(b) 2.0×10^{-4} moles/lit
(c) 1.0×10^{-5} moles/lit
(d) 1.2×10^{-12} moles/lit
9. Solubility product of $AgCl$ is 1.44×10^{-4} moles/lit at 373K. The solubility of the $AgCl$ in boiling of water will be....
(a) 0.72×10^{-4} moles/lit
(b) 0.72×10^{-2} moles/lit
(c) 1.20×10^{-2} moles/lit
(d) 1.20×10^{-4} moles/lit
10. Precipitation takes place when the ionic product....
(a) Is zero
(b) Exceeds the solubility product
(c) Equal to solubility product
(d) Is less than the solubility product
11. The electromotive force is expressed in the unit of
- (a) Ohm
(b) Volt
(c) Ampere
(d) Coulomb
12. If the cell can provide electrical energy to an external system, it is called as
(a) Galvanic cell
(b) Electrolytic cell
(c) Concentration cell
(d) Chemical cell
13. The current per unit area of electrode surface is known as....
(a) Current efficiency
(b) Current yield
(c) Current density
(d) None of these

Analytical Chemistry Questions And Answers

CH Cherryholmes



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