

## Analytical Chemistry MCQs: TYBSC SEM III

8. The solubility product  $K_{sp}$  of sparingly soluble  $Ag_2CrO_4$  is  $4 \times 10^{-12}$ . The molar solubility of the salt is
- (a)  $1.0 \times 10^{-4}$  moles/lit
  - (b)  $2.0 \times 10^{-4}$  moles/lit
  - (c)  $1.0 \times 10^{-5}$  moles/lit
  - (d)  $1.2 \times 10^{-12}$  moles/lit
9. Solubility product of  $AgCl$  is  $1.44 \times 10^{-4}$  moles/lit at 373K. The solubility of the  $AgCl$  in boiling of water will be....
- (a)  $0.72 \times 10^{-4}$  moles/lit
  - (b)  $0.72 \times 10^{-2}$  moles/lit
  - (c)  $1.20 \times 10^{-2}$  moles/lit
  - (d)  $1.20 \times 10^{-4}$  moles/lit
10. Precipitation takes place when the ionic product....
- (a) Is zero
  - (b) Exceeds the solubility product
  - (c) Equal to solubility product
  - (d) Is less than the solubility product
11. The electromotive force is expressed in the unit of .....
- (a) Ohm
  - (b) Volt
  - (c) Ampere
  - (d) Coulomb
12. If the cell can provide electrical energy to an external system, it is called as
- (a) Galvanic cell
  - (b) Electrolytic cell
  - (c) Concentration cell
  - (d) Chemical cell
13. The current per unit area of electrode surface is known as....
- (a) Current efficiency
  - (b) Current yield
  - (c) Current density
  - (d) None of these

# Analytical Chemistry Mcqs

**Didier Musso**



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