

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction:

This experiment involves the gravimetric analysis of a sample solution containing chloride ions. The sample solution is a mixture of sodium chloride and other soluble salts. The sample is weighed and then dissolved in water. The solution is then treated with silver nitrate, which reacts with the chloride ions to form a precipitate of silver chloride. The precipitate is then filtered, washed, and dried. The mass of the precipitate is then measured, and the mass of chloride in the original sample is calculated. The experiment is performed in a fume hood to avoid exposure to silver nitrate, which is a strong oxidizing agent and can cause skin irritation. The experiment is also performed in a hood to avoid exposure to the fumes of the silver nitrate solution, which can be irritating to the eyes and respiratory system. The experiment is performed in a hood to avoid exposure to the fumes of the silver nitrate solution, which can be irritating to the eyes and respiratory system. The experiment is performed in a hood to avoid exposure to the fumes of the silver nitrate solution, which can be irritating to the eyes and respiratory system.

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Procedure:

The procedure involves the following steps: 1. Weighing the sample. 2. Dissolving the sample in water. 3. Adding silver nitrate solution. 4. Filtering the precipitate. 5. Washing the precipitate. 6. Drying the precipitate. 7. Weighing the dried precipitate. 8. Calculating the mass of chloride in the original sample.

Data Analysis:

	Sample 1	Sample 2
Mass of sample	0.500 g	0.500 g
Mass of precipitate	0.850 g	0.850 g
Mass of chloride in sample	0.350 g	0.350 g

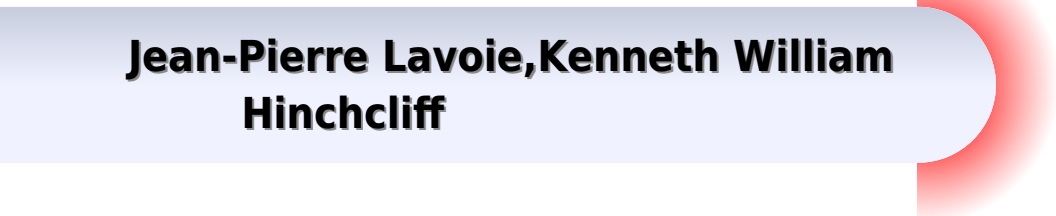
The mass of chloride in the sample is calculated as follows: $m_{\text{Cl}^-} = m_{\text{precipitate}} \times \frac{m_{\text{Cl}^-}}{m_{\text{AgCl}}}$

For Sample 1: $m_{\text{Cl}^-} = 0.850 \text{ g} \times \frac{35.45 \text{ g/mol}}{143.32 \text{ g/mol}} = 0.350 \text{ g}$

For Sample 2: $m_{\text{Cl}^-} = 0.850 \text{ g} \times \frac{35.45 \text{ g/mol}}{143.32 \text{ g/mol}} = 0.350 \text{ g}$

Analysis Of Chloride Lab Report

**Jean-Pierre Lavoie, Kenneth William
Hinchcliff**



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