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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	H-H	O-H
Bond energy in kJ/mol	436	463

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an O-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$(4 \times 436) + 3(498)$$

$$4360 \quad \text{kJ}$$
- Calculate the energy released when bonds form in the products.

$$\text{Bond broken} - \text{Bond formed} = +90$$

$$4360 - x = +90$$

$$x = 4360 - 90$$

$$x = 4270 \quad \text{kJ}$$
- Calculate the energy of an O-H bond.

$$6 \times (\text{O-H}) = 4270$$

$$\text{O-H} = \frac{4270}{6}$$

$$711.67 \quad \text{kJ/mol}$$

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Ovarian cancer is an umbrella term for a collection of distinct disease entities histotypes Until now the majority of research has focussed on the most common histotype high grade serous ovarian carcinoma HGSOC which accounts for 70% of cases The less common histotypes including endometrioid clear cell mucinous low grade serous carcinosarcoma and non epithelial histotypes have received far less research attention Accordingly while major advances in our understanding of HGSOC have led to molecularly directed therapies that improve patient outcomes progress in less common histotypes has lagged behind Notably many of the uncommon histotypes demonstrate resistance to conventional chemotherapy regimens and their inherent biological differences suggest most are unlikely to respond to emerging molecular therapeutics designed to target the biology of common ovarian cancer types *Biomedical Index to PHS-supported Research* ,1990

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