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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Table 3.1 shows some bond energies.

Table 3.1

bond	H-H	N-H
bond energy in kJ/mol	435	413

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an N-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$(4 \times 435) + 1(413)$$

$$4 \times 435 + 1(413)$$

$$2150$$
- Calculate the energy released when bonds form in the products.

$$\text{bond broken} - \text{bond formed} = -92$$

$$2150 - x = -92$$

$$x = 2150 + 92$$
- Calculate the energy of an N-H bond.

$$6 \times \text{N-H} = 2148$$

$$\text{N-H} = \frac{2148}{6}$$

$$358$$

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Xiaolong Qi



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