



Cambridge Assessment
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IGCSE A* MODEL ANSWER

0620 CHEMISTRY

May/June 2024

Paper 42



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Fig. 3.1

(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1 shows some bond energies.

bond	Energy	Unit
bond energy in kJ/mol	440	kJ/mol

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of one C-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$\begin{aligned}
 & (2000) + 3(440) \\
 & 440 + 3(440)
 \end{aligned}
 \quad \underline{3160} \text{ kJ/mol}$$
- Calculate the energy released when bonds form in the products.

$$\begin{aligned}
 \text{bond broken} - \text{bond formed} & = -90 \\
 3160 - x & = -90 \\
 x & = 3250 + 90
 \end{aligned}
 \quad \underline{3340} \text{ kJ/mol}$$
- Calculate the energy of one C-H bond.

$$\begin{aligned}
 6 \times \text{C-H} & = 3340 \\
 \text{C-H} & = \frac{3340}{6}
 \end{aligned}
 \quad \underline{556.67} \text{ kJ/mol}$$

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