

1999 Advanced Placement Chemistry Exam

Section I: Multiple Choice

Time — 1 hour and 30 minutes

NO CALCULATORS MAY BE USED WITH SECTION I.

Note: For all questions, assume that the temperature is 298 K, the pressure is 1.00 atmosphere, and solutions are aqueous unless otherwise specified.

Throughout the test the following symbols have the definitions specified unless otherwise noted.

T = temperature
 P = pressure
 V = volume
 S = entropy
 H = enthalpy
 G = free energy
 R = molar gas constant
 n = number of moles
 M = molar

m = molal
 L, mL = liter(s), milliliter(s)
 g = gram(s)
 nm = nanometer(s)
 atm = atmosphere(s)
 J, kJ = joule(s), kilojoule(s)
 V = volt(s)
 mol = mole(s)

Part A

Directions: Each set of lettered choices below refers to the numbered statements immediately following it. Select the one lettered choice that best fits each statement and then fill in the corresponding oval on the answer sheet. A choice may be used once, more than once, or not at all in each set.

Questions 1–4 refer to the following types of energy.

- (A) Activation energy
- (B) Free energy
- (C) Ionization energy
- (D) Kinetic energy
- (E) Lattice energy

1. The energy required to convert a ground-state atom in the gas phase to a gaseous positive ion
2. The energy change that occurs in the conversion of an ionic solid to widely separated gaseous ions
3. The energy in a chemical or physical change that is available to do useful work
4. The energy required to form the transition state in a chemical reaction

Questions 5–8 refer to atoms for which the occupied atomic orbitals are shown below.

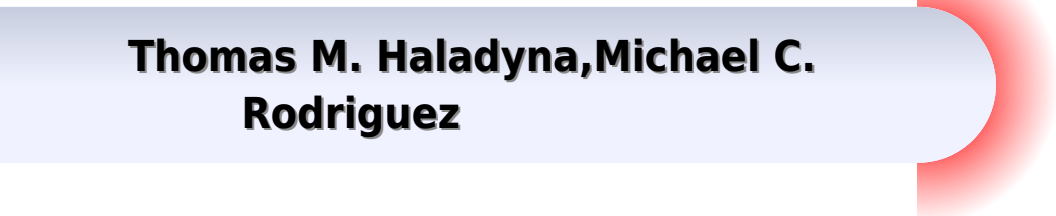
- (A) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}}$
- (B) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}}$
- (C) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}} 2p \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$
- (D) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}} 2p \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$
- (E) $[Ar] 4s \underline{\hspace{1cm}} 3d \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$

5. Represents an atom that is chemically unreactive
6. Represents an atom in an excited state
7. Represents an atom that has four valence electrons
8. Represents an atom of a transition metal

Questions 9–12 refer to aqueous solutions containing 1:1 mole ratios of the following pairs of substances. Assume all concentrations are 1 M .

2008 Ap Chemistry Exam Multiple Choice Answers

**Thomas M. Haladyna, Michael C.
Rodriguez**



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Developing and Validating Test Items Thomas M. Haladyna, Michael C. Rodriguez, 2013-07-18 Since test items are the building blocks of any test learning how to develop and validate test items has always been critical to the teaching learning process As they grow in importance and use testing programs increasingly supplement the use of selected response multiple choice items with constructed response formats This trend is expected to continue As a result a new item writing book is needed one that provides comprehensive coverage of both types of items and of the validity theory underlying them This book is an outgrowth of the author's previous book *Developing and Validating Multiple Choice Test Items* 3e Haladyna 2004 That book achieved distinction as the leading source of guidance on creating and validating selected response test items Like its predecessor the content of this new book is based on both an extensive review of the literature and on its author's long experience in the testing field It is very timely in this era of burgeoning testing programs especially when these items are delivered in a computer based environment Key features include Comprehensive and Flexible No other book so thoroughly covers the field of test item development and its various applications Focus on Validity Validity the most important consideration in testing is stressed throughout and is based on the Standards for Educational and Psychological Testing currently under revision by AERA APA and NCME Illustrative Examples The book presents various selected and constructed response formats and uses many examples to illustrate correct and incorrect ways of writing items Strategies for training item writers and developing large numbers of items using algorithms and other item generating methods are also presented Based on Theory and Research A comprehensive review and synthesis of existing research runs throughout the book and complements the expertise of its authors

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The Publishers Weekly, 2008

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through the college application process with tips on entrance essays financial aid forms and campus visits **The Science Teacher** ,2009 *Oswaal JEE Advance 10 Mock Test Papers (Paper-1 & Paper-2) Physics, Chemistry, Mathematics (For 2023 Exam)* Oswaal Editorial Board,2023-02-27 Benefits of book which distinguish it from others 100% Updated With 2022 Paper 1 2 Fully Solved Extensive Practice With 10 Sample Question Papers 100% Exam Readiness With highly Probable SQPs with a success rate of more than 87% in 2022 Crisp Revision With Mind Maps Mnemonics Appendix Valuable Exam Insights With Subjective Trend Analysis

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