

Worksheet 1A: Remembering General Chemistry: Electronic Structure and Bonding



CHE.231R. _____

Team # _____

Names: _____ ID : m _____

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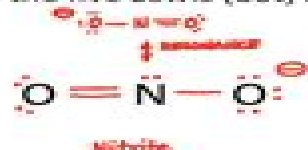
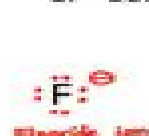
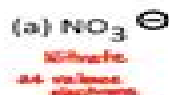
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Fill in the following blanks with the atomic symbol of the element which best fits the following description:

(a) The element with the electronic configuration $1s^2 2s^2 2p^6 3s^2 3p^2$ Si (silicon); (b) the element with twelve protons in its nucleus Mg (magnesium); (c) the second row element that is unchanged when it forms three bonds and has two nonbonding electrons N (nitrogen); (d) the most electronegative element in the third row of the periodic table Cl (chlorine); (e) the atom below boron Al (aluminum); the atom to the right of carbon N (nitrogen).

Fill in the blanks: Atoms with the same number of protons but different numbers of neutrons are called isotopes. Diazomethane is a pale yellow and explosive gas, with the molecular formula of CH_2N_2 and 16 (sixteen) number of valence electrons. Oxygen atoms have a positive / $\oplus$ formal charge when bonded to three atoms. The shape of the ammonium ion $\oplus\text{NH}_4$ is said to be tetrahedral, whereas the shape of carbon dioxide CO_2 is linear.

1. Complete the five Lewis (dot) structures so that each of the following will be **negatively charged** ions:2. From the molecular formulas, draw *reasonable* Lewis structures which clearly illustrate all **nonbonding electrons** and any **formal charges** where necessary:

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