(d)	Circle and label on the graph the point you consider to be the most anomalous. Do not circle or label any other point. If it is assumed that the analysis was carried out accurately, suggest a reason why the point might be anomalous.
	[2]

(e) By drawing an appropriate construction line on your graphs, determine the mole fraction of CCl₄ in the vapour which is in equilibrium with a liquid with a mole fraction of 0.500 CCl₄.

(f) The temperatures were measured using a thermometer calibrated in 0.1 °C graduations. If the thermometer had only been calibrated in 1.0 °C graduations, calculate the percentage errors which would result from the determination of the boiling points of each of the two pure liquids.

[2]

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