

Chemical Kinetics

- 1) Molecular hydrogen can be made from methane gas by the reaction below. How is the rate of disappearance of CH_4 related to the rate of appearance of H_2 ?



- A) _____
 B) _____
 C) _____
 D) none of these

Answer: B

- 2) For the reaction shown below, what is the relationship between the rate of formation of Br_2 and the rate of reaction of Br^- over the same time period?



- A) rate of formation of $\text{Br}_2 = 1/2 \times$ rate of reaction of Br^-
 B) rate of formation of $\text{Br}_2 = 1/4 \times$ rate of reaction of Br^-
 C) rate of formation of $\text{Br}_2 = 2 \times$ rate of reaction of Br^-
 D) rate of formation of $\text{Br}_2 = 2 \times$ rate of reaction of Br^-

Answer: A

- 3) The reaction that occurs in a Breathalyzer, a device used to determine the alcohol level in a person's bloodstream, is given below. If the rate of appearance of $\text{Cr}_2(\text{SO}_4)_3$ is 1.34 mol/min at a particular moment, what is the rate of disappearance of $\text{C}_2\text{H}_5\text{O}$ at that moment?



- A) 0.413 mol/min
 B) 0.826 mol/min
 C) 1.86 mol/min
 D) 3.72 mol/min

Answer: C

Ap Chem Multiple Choice Answers

John T. Moore, Richard H. Langley



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