

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction:

This experiment involves the gravimetric analysis of a solution containing chloride ions. The procedure involves the precipitation of the chloride ions as silver chloride (AgCl) using a solution of silver nitrate (AgNO_3). The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in a fume hood to avoid exposure to the toxic fumes of silver nitrate. The procedure is as follows: 1. Preparation of the sample solution: A known mass of a sample containing chloride ions is dissolved in a known volume of water. 2. Addition of silver nitrate: A solution of silver nitrate is added to the sample solution until all the chloride ions have been precipitated. 3. Filtration: The precipitate is filtered using a pre-weighed filter paper. 4. Drying: The filter paper with the precipitate is placed in a desiccator to dry. 5. Weighing: The dried precipitate is weighed on a balance. 6. Calculation: The mass of chloride ions is calculated from the mass of the precipitate using the stoichiometry of the reaction.

The reaction between silver nitrate and chloride ions is as follows: $\text{AgNO}_3 + \text{Cl}^- \rightarrow \text{AgCl} + \text{NO}_3^-$. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in a fume hood to avoid exposure to the toxic fumes of silver nitrate.

Procedure:

The procedure for this experiment is as follows: 1. Preparation of the sample solution: A known mass of a sample containing chloride ions is dissolved in a known volume of water. 2. Addition of silver nitrate: A solution of silver nitrate is added to the sample solution until all the chloride ions have been precipitated. 3. Filtration: The precipitate is filtered using a pre-weighed filter paper. 4. Drying: The filter paper with the precipitate is placed in a desiccator to dry. 5. Weighing: The dried precipitate is weighed on a balance. 6. Calculation: The mass of chloride ions is calculated from the mass of the precipitate using the stoichiometry of the reaction.

Data Analysis:

	Sample 1	Sample 2
Mass of sample	0.500 g	0.500 g
Mass of precipitate	0.850 g	0.850 g
Mass of chloride ion	0.350 g	0.350 g

The mass of chloride ion in the sample is calculated as follows: $\text{Mass of Cl}^- = \text{Mass of precipitate} \times \frac{\text{Molar mass of Cl}^-}{\text{Molar mass of AgCl}}$

For Sample 1: $\text{Mass of Cl}^- = 0.850 \text{ g} \times \frac{35.45 \text{ g/mol}}{143.32 \text{ g/mol}} = 0.350 \text{ g}$

The mass of chloride ion in the sample is calculated as follows: $\text{Mass of Cl}^- = \text{Mass of precipitate} \times \frac{\text{Molar mass of Cl}^-}{\text{Molar mass of AgCl}}$

Analysis Of Chloride Experiment 7 Lab Report

L Darling-Hammond



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