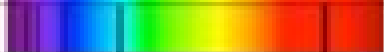


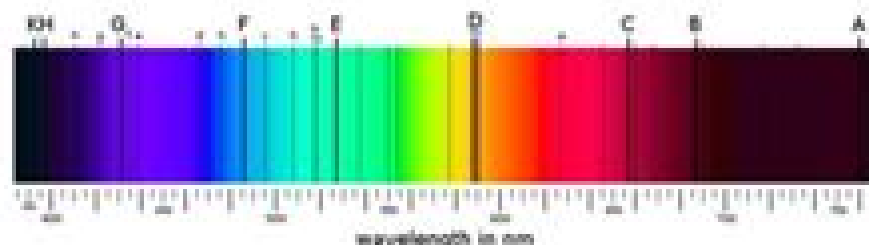
# Spectral Analysis

Worksheet

## Types of spectra:

| Type of spectrum          | Photographic example                                                              |
|---------------------------|-----------------------------------------------------------------------------------|
| Continuous (or continuum) |  |
| Absorption (dark line)    |  |
| Emission (bright line)    |  |

Here is a spectrum with many of the known spectral lines.



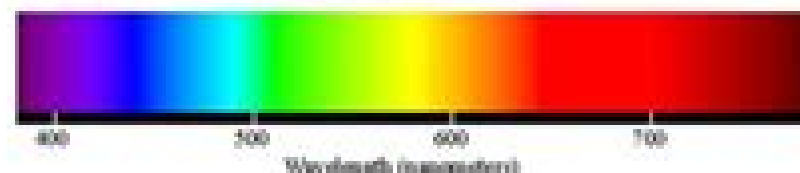
1. What colour is a photon of light with wavelength

- a. 450 nm \_\_\_\_\_ b. 700 nm \_\_\_\_\_  
c. 523 nm \_\_\_\_\_ d. 560 nm \_\_\_\_\_

2. What colour is the emission of hydrogen-alpha, H $\alpha$ ?

3. Draw the main spectrum lines for a star containing Iron and Hydrogen.

(note: Some elements have more than one absorption line)



4. Draw the Emission Lines on a spectrum for a Magnesium lamp. Insert a wavelength values



5. A Nebula of mainly Hydrogen and Mercury (unlikely) is between a star and the earth.

a. What happens to the light as it passes through the nebula

\_\_\_\_\_  
\_\_\_\_\_

b. The Star is mainly hydrogen and Helium. Complete the spectrum



| Designation | Element    | Wavelength (nm) | Designation | Element    | Wavelength (nm) |
|-------------|------------|-----------------|-------------|------------|-----------------|
| y           | O $_2$     | 844.663         | a           | Fe         | 485.763         |
| Z           | O $_2$     | 844.694         | P           | H $\beta$  | 486.134         |
| A           | O $_2$     | 777.419         | d           | Fe         | 486.834         |
| B           | O $_2$     | 686.719         | e           | Fe         | 438.335         |
| C           | H $\alpha$ | 656.281         | G           | H $\gamma$ | 434.047         |
| e           | O $_2$     | 627.641         | G           | Fe         | 430.790         |
| D $_1$      | Na         | 589.592         | G           | Ca         | 430.774         |
| D $_2$      | Na         | 589.595         | h           | H $\delta$ | 430.175         |
| D $_1$ or d | H $\alpha$ | 587.5638        | H           | Ca         | 396.847         |
| e           | H $\alpha$ | 544.813         | K           | Ca         | 395.348         |
| E $_1$      | Fe         | 527.659         | L           | Fe         | 382.044         |
| h $_1$      | Mg         | 518.342         | N           | Fe         | 378.121         |
| h $_2$      | Mg         | 518.379         | P           | Fe         | 376.332         |
| h $_3$      | Fe         | 516.891         | T           | Fe         | 362.108         |
| h $_4$      | Fe         | 516.891         | i           | Fe         | 359.444         |
| h $_5$      | Mg         | 516.733         |             |            |                 |

# Analysis Of Spectral Lines Answers

**R.G. Athay**



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