

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction

This experiment involves the gravimetric analysis of a solution containing chloride ions. The procedure involves the precipitation of the chloride ions as silver chloride (AgCl) using a solution of silver nitrate (AgNO_3). The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in a fume hood to avoid exposure to the toxic fumes of silver nitrate. The procedure is as follows: 1. Preparation of the sample solution: A known mass of a sample containing chloride ions is dissolved in a known volume of water. 2. Addition of silver nitrate: A solution of silver nitrate is added to the sample solution until all the chloride ions have been precipitated as silver chloride. 3. Filtration: The precipitate is filtered from the solution using a filter paper. 4. Drying: The precipitate is dried in a desiccator to remove any water. 5. Weighing: The dried precipitate is weighed on a balance. 6. Calculation: The mass of the precipitate is used to calculate the mass of chloride ions in the original solution.

The reaction between silver nitrate and chloride ions is as follows: $\text{AgNO}_3 + \text{Cl}^- \rightarrow \text{AgCl} + \text{NO}_3^-$. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The calculation is as follows: $\text{Mass of Cl}^- = \frac{\text{Mass of AgCl}}{\text{Molar mass of AgCl}} \times \text{Molar mass of Cl}^-$.

Procedure

The procedure for this experiment is as follows: 1. Preparation of the sample solution: A known mass of a sample containing chloride ions is dissolved in a known volume of water. 2. Addition of silver nitrate: A solution of silver nitrate is added to the sample solution until all the chloride ions have been precipitated as silver chloride. 3. Filtration: The precipitate is filtered from the solution using a filter paper. 4. Drying: The precipitate is dried in a desiccator to remove any water. 5. Weighing: The dried precipitate is weighed on a balance. 6. Calculation: The mass of the precipitate is used to calculate the mass of chloride ions in the original solution.

Data Analysis

	Sample 1	Sample 2
Mass of sample	0.500 g	0.500 g
Mass of precipitate	0.850 g	0.850 g
Mass of chloride ion	0.350 g	0.350 g

The mass of the precipitate is used to calculate the mass of chloride ions in the original solution.

$$\text{Mass of Cl}^- = \frac{\text{Mass of AgCl}}{\text{Molar mass of AgCl}} \times \text{Molar mass of Cl}^-$$

The mass of the precipitate is used to calculate the mass of chloride ions in the original solution.

Analysis Of Chloride Lab Report

Y. V. Singh



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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the iron metal dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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stirring the solution adding water then removing water adding acetone heating the acetone on low

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web stoichiometry lab iron and copper sulfate sample data data and observations mass of copper ii sulfate pentahydrate 3.27 g of g mass of filter paper mass of filter paper and dry copper mass of dry copper 1.28 g

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