

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction:

This experiment involves the gravimetric analysis of a sample solution containing chloride ions. The sample solution is a mixture of sodium chloride and other soluble salts. The sample is weighed and then dissolved in water. The solution is then treated with silver nitrate, which reacts with the chloride ions to form a precipitate of silver chloride. The precipitate is then filtered, washed, and dried. The mass of the dried precipitate is then measured. The mass of the chloride ions in the original sample can be calculated from the mass of the dried precipitate.

The reaction between silver nitrate and sodium chloride is as follows:

$$\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$$

Procedure:

The procedure for this experiment is as follows:

1. Weigh a sample of the solution (approximately 0.5 g).
2. Dissolve the sample in water in a beaker.
3. Add a few drops of nitric acid to the solution.
4. Add a solution of silver nitrate to the solution until a precipitate forms.
5. Filter the precipitate using a filter paper.
6. Wash the precipitate with distilled water.
7. Dry the precipitate in an oven at 100°C.
8. Weigh the dried precipitate.

Data Analysis:

	Sample 1	Sample 2
Mass of sample	0.50 g	0.50 g
Mass of precipitate	0.85 g	0.85 g
Mass of chloride ion	0.42 g	0.42 g

The mass of the chloride ion in the sample can be calculated from the mass of the dried precipitate:

$$\text{Mass of Cl}^- = \text{Mass of AgCl} \times \frac{\text{Molar mass of Cl}^-}{\text{Molar mass of AgCl}}$$

The mass of the chloride ion in the sample can be calculated from the mass of the dried precipitate:

Analysis Of Chloride Experiment 7 Lab Report

Jessica J Manson



Analysis Of Chloride Experiment 7 Lab Report:

Embracing the Melody of Term: An Emotional Symphony within **Analysis Of Chloride Experiment 7 Lab Report**

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